

## Exam I Study Guide

- **Physical Changes vs. Chemical Changes (Can you undo it?)**
  - Physical changes result in changes of appearance, but not composition.
    - Melting, evaporation, freezing, etc.
  - Chemical changes result in changes of composition of matter. (Chemical reaction)
    - Burning, mixing, dissolving, etc.
- **Physical Properties vs. Chemical Properties**
  - Physical properties can be described by observing the matter without a change in the composition.
    - Color, odor, density, temperature, etc.
  - Chemical properties can only be detected by altering the composition.
    - Flammability, oxidization, etc.
- **Atoms vs. Molecules**
  - Atoms are the smallest unit of a given element that carry the properties of that element.
  - Molecules are composed of two or more atoms.
- **Elements vs. Compounds**
  - Elements occur in nature and cannot be broken down to simpler particles through ordinary means.
  - Compounds are the combination of two or more elements.
- **Accuracy vs. Precision**
  - Accuracy refers to the agreement of a measurement with a known or accepted value. (Bullseye)
    - Correctness
  - Precision refers to the agreement of a measurement with other measurements in a given set of measurements. (Missing the bullseye but landing all darts in one place)
    - Repeatability or reproducibility
- **Rules Concerning Significant Figures**
  - All nonzero digits ARE significant.
  - Zeros between nonzero digits ARE significant.
  - Zeros at the beginning of a number are NOT significant; they only indicate the position of the decimal point.
  - Zeros at the end of a number to the right of a decimal point ARE significant.
  - Zeros at the end of a number with no decimal point are ASSUMED to be NOT significant, UNLESS we indicate their significance by placing a line over the last zero that is significant.

# General Chemistry Final Study Guide

**JR Anderson**



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