

**Chem 101 Practice Test: Thermodynamics 3.5 – 3.6, 18.1****Equations**

$$\Delta G = \Delta H - T\Delta S$$

$$\Delta G = \Delta H_{\text{sys}} - T\Delta S_{\text{sys}}$$

$$\Delta G = \Delta H_{\text{sys}}$$

$$\Delta H_{\text{sys}} = \sum \text{Enthalpies of formation}$$

$$\Delta S = \sum S - \sum R$$

**Properties of water**

$$\Delta H_{\text{vap}} = 40.7 \text{ kJ/mol}$$

$$\Delta H_{\text{fus}} = 6.01 \text{ kJ/mol}$$

$$\Delta H_{\text{cond}} = -40.7 \text{ kJ/mol}$$

$$\Delta H_{\text{sub}} = 46.7 \text{ kJ/mol}$$

$$\Delta H_{\text{melt}} = 6.01 \text{ kJ/mol}$$

**Multiple choice**

Multiple choice and multiple-choice questions follow. One or more may be correct.

(1) spontaneous

(2) entropy

(3) Gibbs

(4) enthalpy

(5) heat capacity

(6) specific heat

(7) Gibbs function

(8) heat of vaporization

(9) heat of fusion

- ☒ 1. spontaneity of thermodynamic processes depends on the temperature,  $\Delta G$  of system for  $T=0$ .
- ☒ 2. full work of energy
- ☒ 3. spontaneity of thermodynamic process depends on the temperature,  $\Delta G$  of system for  $T=0$ .
- ☒ 4. thermodynamic processes that have thermodynamic equilibrium during a chemical reaction process
- ☒ 5. balance a system at constant pressure
- ☒ 6. the enthalpy change is used for describing equilibrium
- ☒ 7. the energy relationship is established at the reaction point

**Multiple choice questions**

8. Which of the following is correct for a system?
- ☒ a. A system has a fixed mass and energy.
- ☒ b. A system has a fixed mass and energy.
- ☐ c. A system has a fixed mass and energy.
- ☐ d. A system has a fixed mass and energy.
9. Which of the following is correct for a thermodynamic system?
- ☒ a. enthalpy change
- ☒ b. enthalpy change
- ☒ c. enthalpy change
- ☒ d. enthalpy change
10. In a process that is reversible, the \_\_\_\_\_
- ☒ a. enthalpy change is zero
- ☒ b. enthalpy change is zero
- ☒ c. enthalpy change is zero
- ☒ d. enthalpy change is zero

# Genchem Practice Test Thermochemistry

**Alexander Stone Macnow**



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