

PART I - TRANSPORT PROCESSES: MOMENTUM, HEAT, AND MASS

(1.2-1)

Eq. (1.3-4) $K = 353.2 = ^\circ\text{C} + 273.15$ $^\circ\text{C} = 80.05 = \boxed{80^\circ\text{C}}$

Eq. (1.3-2) $^\circ\text{C} = 80.0 = (1/1.8)(^\circ\text{F} - 32)$ $^\circ\text{F} = \boxed{176^\circ\text{F}}$

Eq. (1.3-3) $^\circ\text{R} = ^\circ\text{F} + 460 = 176 + 460$ $^\circ\text{R} = \boxed{636^\circ\text{R}}$

(1.2-2)

Eq. (1.3-2) $^\circ\text{C} = (1/1.8)(^\circ\text{F} - 32) = (1/1.8)(155 - 32) = \boxed{68.33^\circ\text{C}}$

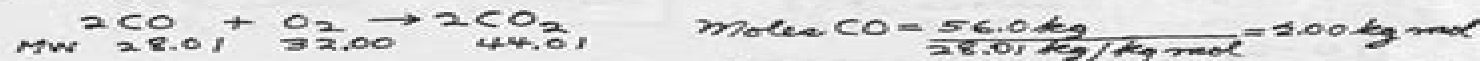
Eq. (1.3-4) $K = ^\circ\text{C} + 273.15 = 68.33 + 273.15 = \boxed{341.5\text{ K}}$

Eq. (1.3-3) $^\circ\text{R} = ^\circ\text{F} + 460 = 155 + 460 = \boxed{615^\circ\text{R}}$

(1.3-1)

$MW(\text{O}_2) = 32.00, MW(\text{N}_2) = 28.02$
 $MW(\text{air}) = 0.21(32.00) + 0.79(28.02) = \boxed{28.9\text{ kg/kg mol}}$

(1.3-2)



Moles CO = $\frac{56.0\text{ kg}}{28.01\text{ kg/kg mol}} = 2.00\text{ kg mol}$

Moles O₂ = 1.00, $1.00(32.00) = \boxed{32.0\text{ kg O}_2}$ $2.00(44.0) = \boxed{88.0\text{ kg CO}_2}$

(1.3-3)

Gas	g	MW	g mol	Mol free
N ₂	20	28.02	0.7138	0.1648 mol free
O ₂	83	32.00	2.5938	0.5990
CO ₂	45	44.01	1.0225	0.2362
Total	148 g		4.3301 g mol	1.0000

Average mol wt = $\frac{148\text{ g}}{4.3301\text{ g mol}} = \boxed{34.2\text{ g/g mol}}$
 $\boxed{34.2\text{ kg/kg mol}}$

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