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රසායන විදනව I සිදුණාගනන්ගන I Chemistry I



පැප දෙකයි මුදුක්තු හක්කිළුවගතවර Two hours

#### Imstructionse

- \* Periodic Table is provided.
- This paper consists of 09 pages.
- · Answer all the questions.
- Use of calculators is not allowed.
- Write your Index Number in the space provided in the answer sheet.
- Follow the instructions given on the back of the answer sheet carefully.
- \* In each of the questions 1 to 50, pick one of the alternatives from (1), (2), (3), (4), (5) which is correct or most appropriate and mark your response on the answer sheet with a cross (x) in accordance with the instructions given on the back of the answer sheet.

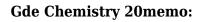
Universal gas constant  $R = 8.314 \text{ J K}^{-1} \text{ mol}^{-1}$ Avogadro constant  $N_a = 6.022 \times 10^{23} \text{ mol}^{-1}$  Planck's constant  $h = 6.626 \times 10^{-34} \text{ J s}$ Velocity of light  $c = 3 \times 10^8 \text{ m s}^{-1}$ 

- I. Which of the following electronic transitions belongs to the visible region of the line spectrum of atomic hydrogen? (n = Principal quantum number)
  - (1)  $n=5 \rightarrow n=3$ (2)  $n=4 \rightarrow n=2$
  - (3)  $n = 1 \rightarrow n = 2$
  - (4)  $n=3 \rightarrow n=1$
  - (5)  $n=2 \rightarrow n=1$
- 2. Select the incorrect statement.
  - (1) Pauli exclusion principle excludes the possibility of the presence of a third electron in an orbital.
  - (2) In a potassium atom, the number of electrons having quantum numbers n (principal quantum number) = 3 and m, (magnetic quantum number) = 0 is 4.
  - (3) The effective nuclear charge felt by a valence electron in nitrogen (N), is greater than the effective nuclear charge felt by a valence electron in carbon (C).
  - (4) Among the ions Na\*, Mg<sup>2\*</sup>, K\* and Ca<sup>2\*</sup>, the two ions closest to each other in size are K\* and Mg<sup>2\*</sup>.
  - (5) Electron gain energy of carbon is negative.
- The increasing order of the second ionization energy (X\*(g) → X²\*(g) + e) of Be, B and O is,
  - (I) Be < B < O
  - (2) Be < O < B (3) B < O < Be
  - (4) B < Be < O
  - (5) O < Be < B
- 4. The shapes of F<sub>4</sub>ClO, FClO, and FClO, are respectively,
  - (1) tetrahedral, trigonal planar and see-saw.
  - (2) square planar, trigonal planar and tetrahedral.
  - (3) see-saw, trigonal pyramidal and square planar.
  - (4) tetrahedral, trigonal pyramidal and see-saw.
  - (5) see-saw, trigonal pyramidal and tetrahedral.

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