



Periodic Table of the Elements

Legend:

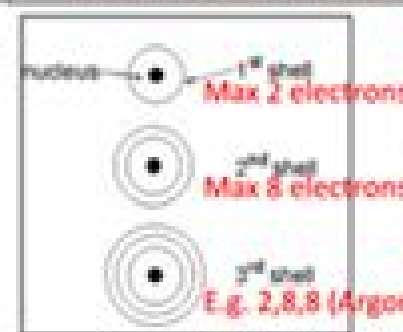
- metals
- non-metals
- metalloids
- transition metals
- alkali metals
- alkali earth metals
- transition metals
- lanthanoids
- actinoids

An atom:
Number of Protons = Number of Electrons
Mass number = Protons + Neutrons

C2 Chemistry

Periodic table arranges elements by
ATOMIC NUMBER (proton number/the small one)

The number of outer shell electrons match the group the element is found in.
E.g. Lithium 2,1 is a group 1 element.

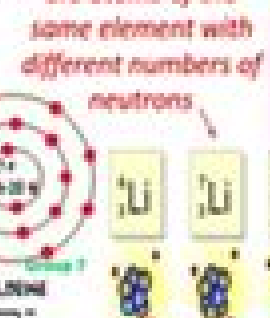
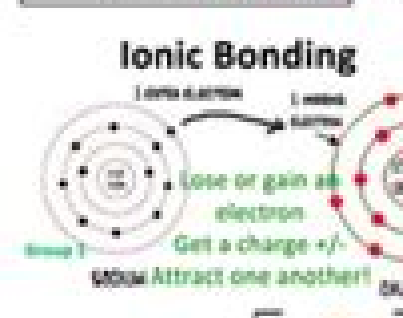
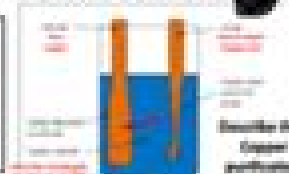


Protons and neutrons both have a mass of 1

ISOTOPES
are atoms of the same element with different numbers of neutrons

OIL RIG

Oxidation is loss (of electrons)
Reduction is gain (of electrons)



When there are 2 possible products
The less reactive forms at the electrode

Electrolysis needs:

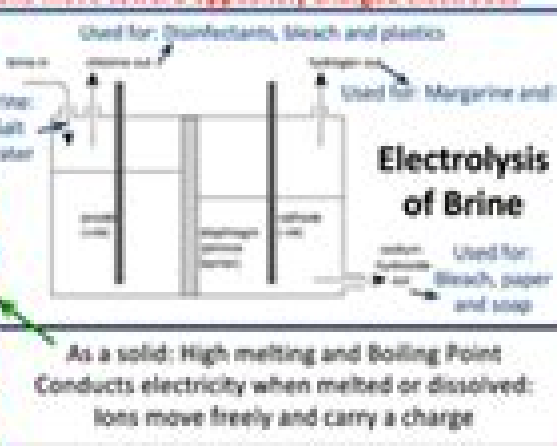
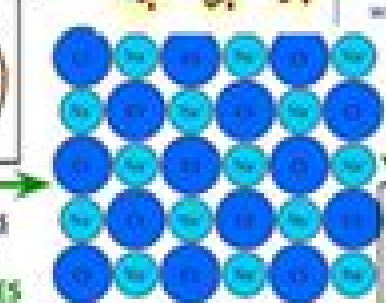
- 2 electrodes
- An electrolyte (ionic or molten substance)
- A power supply

Metal ions only arranged in layers
ELECTROSTATIC ATTRACTION
between
+IONS and ELECTRONS

Metal ions slide over one another making them **MALLEABLE**

Metals conduct **HEAT** well

Giant Lattice Structure
Always metals with non-metals
+ ION and - ION attracted by
STRONG ELECTROSTATIC FORCES



NANO-TECHNOLOGY
Means **REALLY REALLY REALLY** small!!!!

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Guy Waller



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