

Chapter 4 Review Worksheet

Name: _____ Block: _____ Date: _____

1. Which of the following is the smallest particle of an element that can exist by itself?

- A. Ion
- B. Atom**
- C. Molecule
- D. Compound

2. Which of the following correctly matches the subatomic particle with its charge and location in an atom?

| | Subatomic Particle | Location | Charge |
|----------|--------------------|--------------|-----------------|
| A | Proton | Nucleus | Neutral |
| B | Neutron | Nucleus | Positive |
| C | Electron | Shell | Positive |
| D | Electron | Shell | Negative |

3. Which of the following are responsible for bonding?

- A. Nuclei
- B. Protons
- C. Neutrons
- D. Electrons**

4. How do you calculate the number of neutrons in an atom's nucleus?

- A. Atomic number
- B. Mass number - atomic number**
- C. Mass number + atomic number
- D. Number of electrons + Number of protons

5. Which of the following describes a cation?

| | |
|------|--|
| I. | Examples include Ca^{2+} and Al^{3+} |
| II. | A metal atom that has lost electrons |
| III. | Has an equal number of protons and electrons |

- A. I and II only**
- B. I and III only
- C. II and III only
- D. I, II and III

6. Which of the hypothetical elements shown below represents a metal?

- A. Ma
- B. Di
- C. So**
- D. Nh



7. Which of the following does the Bohr model represent?

- A. A neon atom
- B. A sodium atom
- C. A sodium ion**
- D. A fluorine atom



8. Draw Bohr diagrams for the following elements:

a) Nitrogen atom



b) Potassium ion

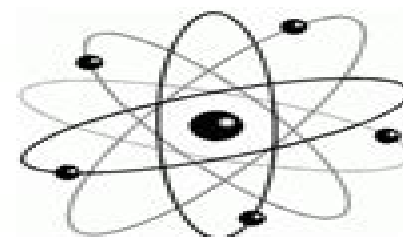


9. Draw Lewis diagrams to show the following chemical bonds:

a) CaCl_2 :



b) CO_2 :



Electrons In Atoms Work Answers

L Towner



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