5

ELECTRONS IN ATOMS

Practice Problems

In your notebook, solve the following problems.

SECTION 5.1 MODELS OF THE ATOM

1. How many sublevels are in the following principal energy levels?

account of the

 $c_{s} \cdot m = 3$

 $b_{x} = 2$

 $d_{\rm e} \approx -4$

e. n = 5f. n = 6

2. How many orbitals are in the following sublevels?

as. Its soublevel

d. 4f sublevel

g. fifth principal energy level

b. 5s sublevel.

es. 7s sublevel

h. 6d sublevel

c. 4d sublevel

Spr. stubskeved

3. What are the types of sublevels and number of orbitals in the following energy levels?

 $a_{m} \cdot m = 0$

 $e_{r} = 3$

 $m_{\rm e} = m_{\rm e} = 5$

 $b_{2} \cdot m = 2$

 $\mathbf{d}_{n} \cdot m = 4.$

SECTION 5.2 ELECTRON ARRANGEMENT IN ATOMS

Write a complete electron configuration of each atom.

a. hydrogen

d. barium

g. kryptom

b. yama-dium

e. bromine

b. arsenic

e. magnesium

f. stalfter

il cadon

SECTION 5.3 PHYSICS AND THE QUANTUM MECHANICAL MODEL

- What is the wavelength of the radiation whose frequency is 5.00 × 10⁴⁵ s⁻⁴?
 In what region of the electromagnetic spectrum is this radiation?
- An inexpensive laser that is available to the public emits light that has a wavelength of 670 nm. What are the color and frequency of the radiation?
- 3. What is the energy of a photon whose frequency is $2.22 \times 10^{14} \, \mathrm{s}^{-1}$?
- What is the frequency of a photon whose energy is 6.00 × 10⁻¹⁵ H.
- Arrange the following types of electromagnetic indiction in order of increasing frequency.

a. infrared

e. visible light

e. unicroverenes

b. cosmic rays

d. radio verves

£ ultraviolet

 Suppose that your favorite AM radio station broadcasts at a frequency of 1600 kHz. What is the wavelength in meters of the radiation from the station?

Electrons In Atoms Answer Key

DJ Losen

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