

5

ELECTRONS IN ATOMS

Practice Problems

In your notebook, solve the following problems.

SECTION 5.1 MODELS OF THE ATOM

- How many sublevels are in the following principal energy levels?

a. $n = 1$	c. $n = 3$	e. $n = 5$
b. $n = 2$	d. $n = 4$	f. $n = 6$
- How many orbitals are in the following sublevels?

a. 1s sublevel	d. 4f sublevel	g. fifth principal energy level
b. 5s sublevel	e. 7s sublevel	h. 6d sublevel
c. 4d sublevel	f. 3p sublevel	
- What are the types of sublevels and number of orbitals in the following energy levels?

a. $n = 1$	c. $n = 3$	e. $n = 5$
b. $n = 2$	d. $n = 4$	

SECTION 5.2 ELECTRON ARRANGEMENT IN ATOMS

- Write a complete electron configuration of each atom.

a. hydrogen	d. barium	g. krypton
b. vanadium	e. bromine	h. arsenic
c. magnesium	f. sulfur	i. radon

SECTION 5.3 PHYSICS AND THE QUANTUM MECHANICAL MODEL

- What is the wavelength of the radiation whose frequency is $5.00 \times 10^{15} \text{ s}^{-1}$? In what region of the electromagnetic spectrum is this radiation?
- An inexpensive laser that is available to the public emits light that has a wavelength of 670 nm. What are the color and frequency of the radiation?
- What is the energy of a photon whose frequency is $2.22 \times 10^{14} \text{ s}^{-1}$?
- What is the frequency of a photon whose energy is $6.00 \times 10^{-19} \text{ J}$?
- Arrange the following types of electromagnetic radiation in order of increasing frequency.

a. infrared	c. visible light	e. microwaves
b. cosmic rays	d. radio waves	f. ultraviolet
- Suppose that your favorite AM radio station broadcasts at a frequency of 1500 kHz. What is the wavelength in meters of the radiation from the station?

Electrons In Atoms Worksheet Answers Chapter 5

Albert A Gayle



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Table of Contents Electrons In Atoms Worksheet Answers Chapter 5

1. Understanding the eBook Electrons In Atoms Worksheet Answers Chapter 5
 - The Rise of Digital Reading Electrons In Atoms Worksheet Answers Chapter 5
 - Advantages of eBooks Over Traditional Books
2. Identifying Electrons In Atoms Worksheet Answers Chapter 5
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electrons In Atoms Worksheet Answers Chapter 5
 - User-Friendly Interface
4. Exploring eBook Recommendations from Electrons In Atoms Worksheet Answers Chapter 5
 - Personalized Recommendations
 - Electrons In Atoms Worksheet Answers Chapter 5 User Reviews and Ratings
 - Electrons In Atoms Worksheet Answers Chapter 5 and Bestseller Lists
5. Accessing Electrons In Atoms Worksheet Answers Chapter 5 Free and Paid eBooks
 - Electrons In Atoms Worksheet Answers Chapter 5 Public Domain eBooks
 - Electrons In Atoms Worksheet Answers Chapter 5 eBook Subscription Services
 - Electrons In Atoms Worksheet Answers Chapter 5 Budget-Friendly Options
6. Navigating Electrons In Atoms Worksheet Answers Chapter 5 eBook Formats

- ePub, PDF, MOBI, and More
 - Electrons In Atoms Worksheet Answers Chapter 5 Compatibility with Devices
 - Electrons In Atoms Worksheet Answers Chapter 5 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electrons In Atoms Worksheet Answers Chapter 5
 - Highlighting and Note-Taking Electrons In Atoms Worksheet Answers Chapter 5
 - Interactive Elements Electrons In Atoms Worksheet Answers Chapter 5
 8. Staying Engaged with Electrons In Atoms Worksheet Answers Chapter 5
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electrons In Atoms Worksheet Answers Chapter 5
 9. Balancing eBooks and Physical Books Electrons In Atoms Worksheet Answers Chapter 5
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electrons In Atoms Worksheet Answers Chapter 5
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Electrons In Atoms Worksheet Answers Chapter 5
 - Setting Reading Goals Electrons In Atoms Worksheet Answers Chapter 5
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Electrons In Atoms Worksheet Answers Chapter 5
 - Fact-Checking eBook Content of Electrons In Atoms Worksheet Answers Chapter 5
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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