

CHAPTER 4 REVIEW

Arrangement of Electrons in Atoms

SECTION 2

SHORT ANSWER Answer the following questions in the space provided.

1. _____ How many quantum numbers are used to describe the properties of electrons in atomic orbitals?
 (a) 1 (b) 2 (c) 3 (d) 4
2. _____ A spherical electron cloud surrounding an atomic nucleus would best represent
 (a) an s-orbital. (b) a p-orbital. (c) a combination of two different p-orbitals. (d) a combination of an s and a p-orbital.
3. _____ How many electrons can an energy level of $n = 4$ hold?
 (a) 32 (b) 24 (c) 8 (d) 6
4. _____ How many electrons can an energy level of $n = 2$ hold?
 (a) 32 (b) 24 (c) 8 (d) 6
5. _____ Compared with an electron for which $n = 2$, an electron for which $n = 4$ has more
 (a) spin. (b) particle nature. (c) energy. (d) wave nature.
6. _____ According to Bohr, which is the point in the figure below where electrons cannot reside?
 (a) point A (b) point B (c) point C (d) point D



7. _____ According to the quantum theory, point D in the above figure represents
 (a) the fixed position of an electron.
 (b) the furthest position from the nucleus that an electron can achieve.
 (c) a position where an electron probably exists.
 (d) a position where an electron cannot exist.

Electrons In Atoms Chapter Test A Answers

Thomas Griffiths



Electrons In Atoms Chapter Test A Answers:

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In a digital era where connections and knowledge reign supreme, the enchanting power of language has are more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Electrons In Atoms Chapter Test A Answers**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound affect our existence. Throughout this critique, we shall delve into the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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