

ANSWER KEY

ATOMS AND ELECTRONS STUDY GUIDE

Learning Goals: (a) (a.i)-(a.ii) Explain a new scientific theory of atomic-scale phenomena (atomic theory) by describing changes in the atomic model over time, including those changes that are consistent with experimental evidence. (b) (b.i)-(b.ii) Explain the relationship between atomic-scale theory and the periodic table, including the periodicity of atomic properties, electronic configurations, ionization energies, electronegativity, and other trends in atomic properties. (c) (c.i)-(c.ii) Explain the relationship between atomic-scale theory and the periodic table, including the periodicity of atomic properties, electronic configurations, ionization energies, electronegativity, and other trends in atomic properties.

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Directions: Match the atomic theory selected on the correct card(s).

Model(s)	Model(s)	Atomic Theory (Select one or more)
(a)	Continuous (solid ball)	A. Everything is made of atoms (that cannot be created, divided or destroyed)
(b)	Discrete (small)	B. Matter is made of small particles (usually in constant motion) that are separated by spaces (intermolecular spaces)
(c)	Discrete (small)	C. Atoms are indivisible (cannot be created)
(d)	Discrete (small)	D. Group Number (Valency) determines
(e)	Discrete (small)	E. Electrons (ions)
(f)	Discrete (small)	F. Atoms are mostly empty space (the subatomic particles move)

Directions: Use the periodic table to answer the following questions:

- Correct** is (a.i)-(a.ii) The number of protons equals the number of electrons.
- Correct** is (b.i)-(b.ii) The atomic number will equal the **Number of protons**.
- Correct** is (c.i)-(c.ii) The number of protons is always greater than the number of electrons (the number of electrons is always negative).

Electrons In Atoms Answer Key Guided

Clemens Wendtner



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