

8 • Electron Configurations & Periodicity

ANSWERS TO STUDY QUESTIONS

- Write the electron configurations of the following elements using the shorthand notation for the noble gas cores.
 - phosphorus [Ne] $3s^2 3p^3$
 - nickel [Ar] $3d^8 4s^2$
 - osmium [Xe] $4f^{14} 5d^6 6s^2$
 - californium [Rn] $5f^{10} 7s^2$
 - titanium [Ar] $3d^2 4s^2$
- Which orbital is filled following these orbitals?
 - 3d is followed by 4p
 - 4s is followed by 3d
 - 5p is followed by 6s
 - 5f is followed by 6d
- How many electrons can be accommodated in
 - a d subshell 10; two in each of 5 orbitals
 - a set of f orbitals 14; two in each of 7 orbitals
 - the $n = 4$ shell 32; $2s^2$
 - the 7s orbital 2
 - a p_x orbital? 2
- What is wrong with the following ground state electron configurations?

<p>a. $\begin{array}{ c c c c } \hline & & & \\ \hline \end{array}$ $\begin{array}{ c } \hline 3d \\ \hline \end{array}$ $\begin{array}{ c c c } \hline & & \\ \hline \end{array}$ $\begin{array}{ c } \hline 4s \\ \hline \end{array}$</p> <p>b. $\begin{array}{ c c c c c } \hline \uparrow\downarrow & \uparrow\downarrow & \uparrow\downarrow & \uparrow\downarrow & \uparrow\downarrow \\ \hline \end{array}$ $\begin{array}{ c } \hline 3d \\ \hline \end{array}$ $\begin{array}{ c } \hline 4s \\ \hline \end{array}$ $\begin{array}{ c c c } \hline \uparrow & \uparrow & \uparrow \\ \hline \end{array}$ $\begin{array}{ c } \hline 4p \\ \hline \end{array}$</p> <p>c. $\begin{array}{ c c c c c } \hline \uparrow\downarrow & \uparrow\downarrow & \uparrow\downarrow & \uparrow\downarrow & \uparrow\downarrow \\ \hline \end{array}$ $\begin{array}{ c } \hline 3d \\ \hline \end{array}$ $\begin{array}{ c } \hline 4s \\ \hline \end{array}$ $\begin{array}{ c c c } \hline \uparrow & \uparrow & \uparrow \\ \hline \end{array}$ $\begin{array}{ c } \hline 4p \\ \hline \end{array}$</p> <p>d. $\begin{array}{ c c c c c } \hline \uparrow & \uparrow & \uparrow & \uparrow & \uparrow \\ \hline \end{array}$ $\begin{array}{ c } \hline 3d \\ \hline \end{array}$ $\begin{array}{ c } \hline 4s \\ \hline \end{array}$ $\begin{array}{ c c c } \hline & & \\ \hline \end{array}$ $\begin{array}{ c } \hline 4p \\ \hline \end{array}$</p> <p>e. $\begin{array}{ c c c c c } \hline \uparrow\downarrow & \uparrow\downarrow & \uparrow\downarrow & \uparrow\downarrow & \uparrow\downarrow \\ \hline \end{array}$ $\begin{array}{ c } \hline 3d \\ \hline \end{array}$ $\begin{array}{ c } \hline 4s \\ \hline \end{array}$ $\begin{array}{ c c c } \hline \uparrow\downarrow & & \\ \hline \end{array}$ $\begin{array}{ c } \hline 4p \\ \hline \end{array}$</p>	<p>a. OK, 4s is filled before 3d.</p> <p>b. 4s should be filled before 4p subshell.</p> <p>c. 4s cannot have 2 electrons with same spin.</p> <p>d. OK, stability of half-filled 3d subshell.</p> <p>e. 4p should first fill one orbital at a time.</p>
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- How many unpaired electrons are there in
 - a nitrogen atom 3 unpaired electrons ($2s^2 2p^3$)
 - an iodine atom 1 unpaired electrons ($5s^2 5p^5$)
 - a nickel (II) cation 2 unpaired electrons ($3d^8$)
 - an oxide ion no unpaired electrons (closed shell [Ne])

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Table of Contents Electronic Configuration And Periodicity Packet Answers

1. Understanding the eBook Electronic Configuration And Periodicity Packet Answers
 - The Rise of Digital Reading Electronic Configuration And Periodicity Packet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Electronic Configuration And Periodicity Packet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electronic Configuration And Periodicity Packet Answers
 - User-Friendly Interface

4. Exploring eBook Recommendations from Electronic Configuration And Periodicity Packet Answers
 - Personalized Recommendations
 - Electronic Configuration And Periodicity Packet Answers User Reviews and Ratings
 - Electronic Configuration And Periodicity Packet Answers and Bestseller Lists
5. Accessing Electronic Configuration And Periodicity Packet Answers Free and Paid eBooks
 - Electronic Configuration And Periodicity Packet Answers Public Domain eBooks
 - Electronic Configuration And Periodicity Packet Answers eBook Subscription Services
 - Electronic Configuration And Periodicity Packet Answers Budget-Friendly Options
6. Navigating Electronic Configuration And Periodicity Packet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Electronic Configuration And Periodicity Packet Answers Compatibility with Devices
 - Electronic Configuration And Periodicity Packet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electronic Configuration And Periodicity Packet Answers
 - Highlighting and Note-Taking Electronic Configuration And Periodicity Packet Answers
 - Interactive Elements Electronic Configuration And Periodicity Packet Answers
8. Staying Engaged with Electronic Configuration And Periodicity Packet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electronic Configuration And Periodicity Packet Answers
9. Balancing eBooks and Physical Books Electronic Configuration And Periodicity Packet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electronic Configuration And Periodicity Packet Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Electronic Configuration And Periodicity Packet Answers
 - Setting Reading Goals Electronic Configuration And Periodicity Packet Answers
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Electronic Configuration And Periodicity Packet Answers
 - Fact-Checking eBook Content of Electronic Configuration And Periodicity Packet Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
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