8 • Electron Configurations & Periodicity

WRITING ELECTRON CONFIGURATIONS

For each given element, fill in the orbital diagram and then write the electron configuration for the element.

1.	2.	3.	4.	5.	6.
	, <u>S</u>				S S S S S S S S S S S S S S S S S S S
2000 O	200000 C	2000	2000	2 × × × × × × × × × × × × × × × × × × ×	80808 80808
¹	¹⊠	180	180	'₩	'₩
Element: Ar # of e 's: 16	Element: Mg # of e-'s: 12	Element: N # of e-'s: 7	Element: Li # of e-'s: _3	Element: P # of e 's: 15	Element: Cl # of e ⁻ 's: 17

Write the electron configurations of each of these in long form and short form:

- Ar 152 252 206 352 306
 - AT TNET 352 3pt
- Mg 152 252 206 352
 - Mg [Ne] 352
- N 152 252 2p3 N [He] 252 2p3
- Li 152 25'
- P 152 252 Zpc 352 3p3
 - P [Ne] 35+ 3p3
- CI 152 252 2p6 352 3p5
- CI [Ne] 352 305



Fill in the orbital diagram for the element, Fe, and write the electron configuration of Fe in the long and short form.

Fe 1s2 252 2p 352 3p6 3d6 452

Fe [Ar] 3d 452



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Robert C. Smoot, Smoot, Richard G. Smith, Jack Price

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Table of Contents Electron Configuration Of Ions Practice Answer Key

- 1. Understanding the eBook Electron Configuration Of Ions Practice Answer Key
 - The Rise of Digital Reading Electron Configuration Of Ions Practice Answer Key
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Electron Configuration Of Ions Practice Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electron Configuration Of Ions Practice Answer Key
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Electron Configuration Of Ions Practice Answer Key
 - Personalized Recommendations
 - Electron Configuration Of Ions Practice Answer Key User Reviews and Ratings
 - Electron Configuration Of Ions Practice Answer Key and Bestseller Lists
- 5. Accessing Electron Configuration Of Ions Practice Answer Key Free and Paid eBooks
 - Electron Configuration Of Ions Practice Answer Key Public Domain eBooks
 - Electron Configuration Of Ions Practice Answer Key eBook Subscription Services

- Electron Configuration Of Ions Practice Answer Key Budget-Friendly Options
- 6. Navigating Electron Configuration Of Ions Practice Answer Key eBook Formats
 - o ePub, PDF, MOBI, and More
 - Electron Configuration Of Ions Practice Answer Key Compatibility with Devices
 - Electron Configuration Of Ions Practice Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electron Configuration Of Ions Practice Answer Key
 - Highlighting and Note-Taking Electron Configuration Of Ions Practice Answer Key
 - Interactive Elements Electron Configuration Of Ions Practice Answer Key
- 8. Staying Engaged with Electron Configuration Of Ions Practice Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electron Configuration Of Ions Practice Answer Key
- 9. Balancing eBooks and Physical Books Electron Configuration Of Ions Practice Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electron Configuration Of Ions Practice Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Electron Configuration Of Ions Practice Answer Key
 - Setting Reading Goals Electron Configuration Of Ions Practice Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Electron Configuration Of Ions Practice Answer Key
 - Fact-Checking eBook Content of Electron Configuration Of Ions Practice Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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