

Electron Configuration Lab

Name _____ KEY _____ Period _____ Date _____

Directions: To understand that electrons fill up the LOWEST energy orbitals first, but that they spread out as much as possible in orbitals.

Part 1: Questions

1. Write the Aufbau Principle.

Electrons will fill the lowest energy orbitals first.

2. Which "rule" or "principle" states that electrons spread out as much as possible within a sublevel before they pair up in the same orbital? **Hund's Rule**

3. In this lab, one cup represents an orbital in the atom. How many cups would be needed to represent the "2p" sublevel?

three (3)

4. How many cups would be needed to represent an atom that has the 1s, 2s, 2p, 3s, 3p sublevels _____?

ten (10)

5. In this lab, beans represent the electrons. How many beans can go into a single cup?

two (2)

Part 2:

Part 2a

Electron Configuration	Total # of Electrons	Element Name	Teacher's Initials
$1s^2 2s^2 2p^3 3s^2 3p^3$	15	Phosphorus	

Part 2b

Station	Electron Configuration	Total # of Electrons
Station 1	$1s^2 2s^2 2p^3$	10
Station 2	$1s^2 2s^2 2p^4$	8
Station 3	$1s^2 2s^2 2p^6 3s^2 3p^2 4s^2$	20
Station 4	$1s^2 2s^2 2p^6 3s^2 3p^4$	18
Station 5	$1s^2 2s^2 2p^3$	9
Station 6	$1s^2 2s^2$	4
Station 7	$1s^2 2s^2 2p^3$	8
Station 8	$1s^2 2s^2$	4
Station 9	$1s^2 2s^2 2p^6 3s^2 3p^4 4s^2$	18
Station 10	$1s^2 2s^2 2p^3$	9

Part 2c: Questions

1. Using a Periodic Table, identify the elements at each station and complete the table to the right.

2. Which orbital is lower in energy -- 4s or 3d?

4s

3. Write the electron configuration for Zinc.

$1s^2 2s^2 2p^6 3s^2 3p^4 4s^2 4p^2$

4. Write the electron configuration for Xenon.

$1s^2 2s^2 2p^6 3s^2 3p^4 4s^2 4p^2 5s^2 5p^4 6s^2 6p^4$

5. Write the electron configuration for Bismuth.

$1s^2 2s^2 2p^6 3s^2 3p^4 4s^2 4p^2 5s^2 5p^4 6s^2 6p^4 7s^2 7p^4$

Station	Element Name
Station 1	Neon (Ne)
Station 2	Oxygen (O)
Station 3	Calcium (Ca)
Station 4	Argon (Ar)
Station 5	Boron (B)
Station 6	Beryllium (Be)
Station 7	Carbon (C)
Station 8	Lithium (Li)
Station 9	Potassium (K)
Station 10	Nitrogen (N)

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drops chlorophyll for 1 c dough yellow 1 teaspoon turmeric powder brighter pink 1 2

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