

Electrochemistry-Ch.9 – PRACTICE TEST/EXAM QUESTIONS

Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

- ____ 1. In the reaction
$$2\text{HNO}_3(\text{aq}) + \text{Cu}(\text{s}) + 2\text{H}^+(\text{aq}) \rightarrow 2\text{NO}_2(\text{g}) + \text{Cu}^{2+}(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$$

Which of the following statements is correct?
a. HNO_3 is the reducing agent, and Cu is the oxidizing agent
b. HNO_3 is the reducing agent, and H^+ is the oxidizing agent
c. Cu is the reducing agent, and H^+ is the oxidizing agent
d. Cu is the reducing agent, and HNO_3 is the oxidizing agent
e. H^+ is the reducing agent, and Cu is the oxidizing agent
- ____ 2. The oxidation number of chromium (Cr) in the dichromate ion ($\text{Cr}_2\text{O}_7^{2-}$) is
a. +2
b. +4
c. +6
d. +7
e. -2
- ____ 3. The oxidation number of fluorine (F) in AlF_3 is
a. -3
b. -2
c. -1
d. 0
e. +1
- ____ 4. In $\text{Fe}_2(\text{SO}_4)_3$ the oxidation numbers of Fe , S , and O respectively are
a. +2, +3, -4
b. +3, +6, -2
c. +2, +4, -8
d. +2, +4, -2
e. +4, +5, -3
- ____ 5. Which substance in the following reaction has undergone oxidation?
$$2\text{MnO}_4^- (\text{aq}) + 5\text{C}_2\text{O}_4^{2-} (\text{aq}) + 16\text{H}^+ (\text{aq}) \rightarrow 2\text{Mn}^{2+} (\text{aq}) + 10\text{CO}_2(\text{g}) + 8\text{H}_2\text{O}(\text{l})$$

a. MnO_4^-
b. $\text{C}_2\text{O}_4^{2-}$
c. H^+
d. Mn^{2+}
e. CO_2
- ____ 6. This question consists of an assertion (statement) in the left-hand column and a reason in the right-hand column.

Assertion

Chlorine will cause iodine to form when chlorine is bubbled through a solution of potassium iodide.

Reason

Chlorine is a stronger oxidizing agent than iodine.

Select one of the following.

- if both assertion and reason are true statements, and the reason is a correct explanation of the assertion
- if both assertion and reason are true statements, and the reason is not a correct explanation of the assertion
- if the assertion is true but the reason is false
- if the assertion is false but the reason is a true statement
- if both assertion and reason are false statements

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Lingjun Ying



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