

1) What is meant by 'limiting molar conductivity'?

The molar conductivity of a solution at infinite dilution is called limiting molar conductivity and is represented by the symbol Λ_m .

2) Express the relation between conductivity and molar conductivity of a solution held in a cell. Explain the terms.

Answer:

$$\Lambda_m = \frac{\kappa}{c} = \frac{\text{Conductivity}}{\text{Concentration}}$$

3) Given that the standard electrode potentials (E°) of metals are :

$\text{K}^+/\text{K} = -2.93 \text{ V}$, $\text{Ag}^+/\text{Ag} = 0.80 \text{ V}$, $\text{Cu}^{2+}/\text{Cu} = 0.34 \text{ V}$,

$\text{Mg}^{2+}/\text{Mg} = -2.37 \text{ V}$, $\text{Cr}^{3+}/\text{Cr} = -0.74 \text{ V}$, $\text{Fe}^{2+}/\text{Fe} = -0.44 \text{ V}$.

Arrange these metals in increasing order of their reducing power.



More negative the value of standard electrode potentials of metals is, more will be the reducing power.

4) Two half-reactions of an electrochemical cell are given below :

$\text{MnO}_4^- (\text{aq}) + 8\text{H}^+ (\text{aq}) + 5\text{e}^- \rightarrow \text{Mn}^{2+} (\text{aq}) + 4\text{H}_2\text{O} (\text{l})$, $E^\circ = 1.51 \text{ V}$

$\text{Sn}^{2+} (\text{aq}) \rightarrow \text{Sn}^{4+} (\text{aq}) + 2\text{e}^-$, $E^\circ = + 0.15 \text{ V}$.

Construct the redox reaction equation from the two half-reactions and calculate the cell potential from the standard potentials and predict if the reaction is reactant or product favoured.

The reactions can be represented at anode and at cathode in the following ways :

At anode (oxidation) :



At cathode (reduction) :



Electrochemistry Problems And Answers

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