Electrochemical Cells and Thermodynamics

17 Pre-lab Questions

Before beginning this experiment in the laboratory, you should be able to answer the following questions.

- 1. Define the following terms: faraday, salt bridge, anode, carbode, voltate cell, and electrolytic cell.
- Write a chemical equation for the reaction that occurs in the following cell: Cu [Cu²⁺(aq)] Ag⁺(aq)] Ag.
- Given the following E's, calculate the standard-cell potential for the cell in question 2.

$$Cu^{2+}(aq) + 2q^{-} \longrightarrow Cu(s)$$
 $E^{+} = +0.34 \text{ V}$
 $Ag^{+}(aq) + e^{-} \longrightarrow Ag(s)$ $E^{-} = +0.80 \text{ V}$

4. Calculate the voltage of the following cell:

5. Calculate the cell potential, the equilibrium constant, and the free-energy change for

$$Ca(x) + Mn^{2+}(aq)(1 M) = Ca^{2+}(aq)(1 M) + Mn(x)$$

given the following E* values:

$$Ca^{2+}(aq) + 2e^- \longrightarrow Ca(s)$$
 $E^* = -2.87 \text{ V}$
 $Ma^{2+}(aq) + 2e^- \longrightarrow Ma(s)$ $E^* = -1.18 \text{ V}$

Electrochemical Lab Pre Lab Answers

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