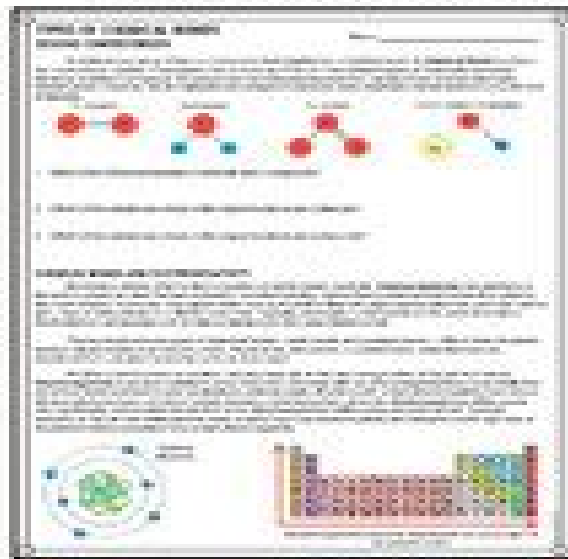


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Types of Chemical Bonds

1. Ionic bonds are formed between a metal and a non-metal. The metal atom loses one or more electrons to become a positively charged ion (cation), and the non-metal atom gains one or more electrons to become a negatively charged ion (anion). The resulting ions are attracted to each other by strong electrostatic forces, forming an ionic bond.

2. Covalent bonds are formed between two non-metal atoms. The atoms share one or more pairs of electrons, forming a covalent bond. The shared electrons are held between the two atoms, pulling them together.

3. Metallic bonds are formed between metal atoms. The metal atoms are arranged in a regular lattice, and the outer electrons are delocalized, forming a 'sea of electrons' that holds the metal ions together.

Diagram showing the formation of ionic, covalent, and metallic bonds.



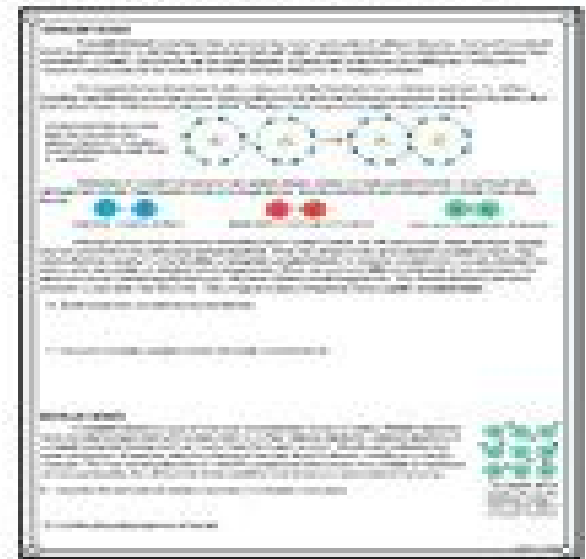
Properties of Ionic Compounds

1. High melting and boiling points: Ionic compounds have strong electrostatic forces between the ions, which require a lot of energy to break.

2. Soluble in water: Many ionic compounds are soluble in water, where the water molecules surround the ions and pull them apart.

3. Conduct electricity: Ionic compounds conduct electricity when molten or dissolved in water, as the ions are free to move.

Diagram showing the structure of an ionic compound lattice.



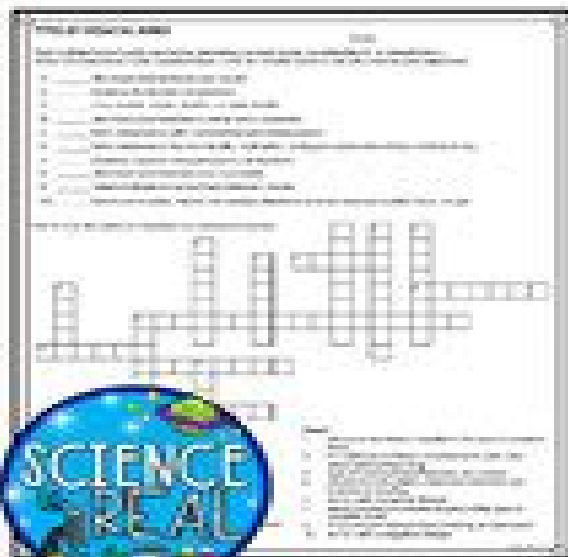
Properties of Covalent Compounds

1. Low melting and boiling points: Covalent compounds have weak intermolecular forces between the molecules, which require less energy to break.

2. Insoluble in water: Most covalent compounds are insoluble in water, as the water molecules cannot break the strong covalent bonds within the molecules.

3. Do not conduct electricity: Covalent compounds do not conduct electricity, as there are no free ions or electrons.

Diagram showing the structure of a covalent compound molecule.



Properties of Metallic Compounds

1. High melting and boiling points: Metallic compounds have strong metallic bonds, which require a lot of energy to break.

2. Soluble in water: Most metallic compounds are insoluble in water, as the water molecules cannot break the strong metallic bonds.

3. Conduct electricity: Metallic compounds conduct electricity, as the delocalized electrons are free to move.

Diagram showing the structure of a metallic compound lattice.

- ✓ IONIC, COVALENT, & METALLIC BONDS
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