

## Chapter 7 Chemical Reactions

**Section 7.1 Describing Reactions**

(pages 192–198)

*This section discusses the use of chemical equations and how to balance them. It also demonstrates the use of calculations in chemistry.*

**Reading Strategy (page 192)**

**Monitoring Your Understanding** Preview the Key Concepts, topic headings, vocabulary, and figures in this section. List two things you expect to learn. After reading, state what you learned about each item you listed. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

What I Expect to Learn	What I Learned
Answers may vary. Possible answers: How to balance chemical equations	Answers may vary. Possible answers: An unbalanced equation can be balanced by changing the coefficients.
How to convert from mass to moles	The mass of a substance can be converted to moles by using the molar mass as a conversion factor.

**Chemical Equations (pages 192–193)**

- Is the following sentence true or false? The new substances formed as a result of a chemical reaction are called products. true
- Circle the letter of each sentence that is a correct interpretation of the chemical equation  $\text{C} + \text{O}_2 \longrightarrow \text{CO}_2$ .
  - Carbon and oxygen react and form carbon monoxide.
  - ☒ Carbon and oxygen react and form carbon dioxide.
  - Carbon dioxide yields carbon and oxygen.
  - ☒ The reaction of carbon and oxygen yields carbon dioxide.
- Is the following sentence true or false? The law of conservation of mass states that mass is neither created nor destroyed in a chemical reaction. true
- Circle the letter of the correct answer. According to the equation  $\text{C} + \text{O}_2 \longrightarrow \text{CO}_2$ , how many carbon atoms react with 14 molecules of oxygen to form 14 molecules of carbon dioxide?
  - 1
  - 7
  - ☒ 14
  - 28
- In the reaction represented by the equation  $\text{C} + \text{O}_2 \longrightarrow \text{CO}_2$ , the mass of carbon dioxide produced equals the total mass of carbon and oxygen that reacted.

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**M Lipman**



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## **Table of Contents Guided Notes Chemical Workbook Answers**

1. Understanding the eBook Guided Notes Chemical Workbook Answers
  - The Rise of Digital Reading Guided Notes Chemical Workbook Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Guided Notes Chemical Workbook Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Guided Notes Chemical Workbook Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Guided Notes Chemical Workbook Answers
  - Personalized Recommendations
  - Guided Notes Chemical Workbook Answers User Reviews and Ratings
  - Guided Notes Chemical Workbook Answers and Bestseller Lists

5. Accessing Guided Notes Chemical Workbook Answers Free and Paid eBooks
  - Guided Notes Chemical Workbook Answers Public Domain eBooks
  - Guided Notes Chemical Workbook Answers eBook Subscription Services
  - Guided Notes Chemical Workbook Answers Budget-Friendly Options
6. Navigating Guided Notes Chemical Workbook Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Guided Notes Chemical Workbook Answers Compatibility with Devices
  - Guided Notes Chemical Workbook Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Guided Notes Chemical Workbook Answers
  - Highlighting and Note-Taking Guided Notes Chemical Workbook Answers
  - Interactive Elements Guided Notes Chemical Workbook Answers
8. Staying Engaged with Guided Notes Chemical Workbook Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Guided Notes Chemical Workbook Answers
9. Balancing eBooks and Physical Books Guided Notes Chemical Workbook Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Guided Notes Chemical Workbook Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Guided Notes Chemical Workbook Answers
  - Setting Reading Goals Guided Notes Chemical Workbook Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Guided Notes Chemical Workbook Answers
  - Fact-Checking eBook Content of Guided Notes Chemical Workbook Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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