

6.9) It is often useful to write the equations for the reactions described:



The reaction for which the heat of reaction is requested is the formation of methane from its constituent elements.



The  $\Delta H$  for this reaction can be calculated by summing the  $\Delta H$ 's for a series of reactions which together add up to the same overall reaction (Hess's Law of Heat Summation).



These equations add up to the desired equation, so the  $\Delta H$  values are added:

$$\Delta H^\circ(\text{CH}_4) = +890. - 394 - 572 \text{ kJ/mol} = -76 \text{ kJ/mol}$$

6.10)  $\Delta H_{\text{vap}} = -924 \text{ kJ}$

6.11) The heat of reaction is the difference between the heat of formation of the products and the heat of formation of the reactants in the balanced equation:

$$\Delta H^\circ = 12 \times (-384) + 8 \times (-242) - 2 \times (+49) = -8278 \text{ kJ.}$$

There are two moles of benzene in the equation, so the heat released for one mole is 1/2 this amount, 3139 kJ/mol.

6.12)  $\Delta H_{\text{vap}} = -95.6 \text{ kJ}$

6.13) a. Heat lost =  $50. \text{ g} \times 0.128 \text{ J/g}^\circ\text{C} \times 68^\circ\text{C} = 260 \text{ J}$

b.  $q(\text{ice melting}) + q(\text{water warming}) = -q(\text{water cooling})$

$$5.0 \text{ g} \times 334 \text{ J/g} + 5.0 \text{ g} \times 4.184 \text{ J/g}^\circ\text{C} (T_f - 0^\circ\text{C}) = -100 \text{ g} \times 4.184 \text{ J/g}^\circ\text{C} (T_f - 70^\circ\text{C})$$

$$1670 \text{ J} + 20.92 T_f = -418.4 T_f + 29,288 \text{ J}$$

$$439.32 T_f = 27,618$$

$$T_f = 62.9 = 63^\circ\text{C}$$

6.14) a.  $q(\text{water}) = -q(\text{coins})$

$$c \times m \times \Delta T = -c \times m \times \Delta T$$

$$4.184 \text{ J/g}^\circ\text{C} \times 200. \text{ g} \times 5.9^\circ\text{C} = -c \times 113.4 \text{ g} \times (-74.1^\circ\text{C})$$

$$\text{Solving for } c \text{ gives: } 0.587 \text{ J/g}^\circ\text{C}$$

b. No, the specific heat capacity of pure gold is  $0.129 \text{ J/g}^\circ\text{C}$ .

6.15)  $\Delta H = 4.18 \text{ J/g}^\circ\text{C} \times 22.4 \text{ K} \times 70. \text{ g} = 6550 \text{ J}$

$$\text{For one mole of ammonium nitrate, } \Delta H = 6550 \times 60.05/20 = 26 \text{ kJ/mol.}$$

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**Bernice G. Segal, Peter S. Shenkin**



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