

STUDY GUIDE: Acids & Bases

- ☐ Define acid and base according to Bronsted-Lowry theory
- ☐ List the general characteristics of acids and bases.
- ☐ Name common binary acids and oxyacids given their chemical formulas
- ☐ Predict the products and balance chemical equations for the reactions with acids involving: reactive metals, metal oxides, metal hydroxides, hydrogen carbonates, and carbonates
- ☐ Explain the process and write balanced equations of neutralization reactions
- ☐ Explain and use the pH scale
- ☐ Use a pH meter and universal indicators to determine the acidity or alkalinity of a substance
- ☐ Define the terms solute, solvent, solution, solubility and molarity
- ☐ Calculate molarity
- ☐ Solve problems involving pH and $[H^+]$
- ☐ Describe how acid rain is formed and the effects of acid rain on humans and ecosystems
- ☐ Determine the number of significant figures in measurements.

1. Bronsted-Lowry Theory

- a. Define acid and base according to **Bronsted-Lowry theory**.
 - b. Give 4 examples of acids and show how they dissociate in water
 - c. How does the dissociation in water relate to the Bronsted-Lowry theory of acids?
 - d. Give at least 2 examples of bases that do not contain hydroxide ions.
2. List the general characteristics of acids and bases.

	Characteristics of Acids	Characteristics of Bases
Taste		
Feel		
pH		
Turn Litmus Paper		
Conductivity?		
Corrosive?		
React with Metals?		

3. Name the following common binary acids and oxyacids given their chemical formulas

- a. HF
 - b. HCN
 - c. H_2SiO_3
 - d. H_3PO_3
 - e. $HClO_2$
 - f. $H_2C_4H_4O_6$
 - g. H_2S
 - h. H_5AsO_4
4. Write the chemical formula for the following acids
- a. Hydrophosphoric acid
 - b. Benzoic acid
 - c. Thiocyanic acid
 - d. Hypochlorous acid
 - e. Bromic acid
 - f. Dichromic acid
 - g. Hydrobromic acid
 - h. Sulfurous acid

Guided And Study Workbook Acids Bases

SJ Ball



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