

Topic 13 Review Key

1. What fraction equals 0.68?

$$\frac{17}{25}$$



The star is underneath what fraction AND what decimal?

Fraction: $\frac{2}{10}$

Decimal: 0.2

3. Kimberly swims $\frac{1}{3}$ mile every day. How many miles will she swim in 6 days?
How many miles will she swim in 7 days?

Kimberly will swim 2 miles in 6 days and $2\frac{1}{3}$ miles in 7 days.

$$\frac{1}{3} \times 6 = \frac{6}{3} = 2 \text{ miles}$$

$$\frac{1}{3} \times 7 = \frac{7}{3} = 2\frac{1}{3} \text{ miles}$$

4. What is 0.08 as a fraction?

a. $\frac{8}{10}$

b. $\frac{1}{8}$

c. $\frac{8}{100}$

d. $\frac{1}{8}$

Envision Math Topic 13 5 Test Answers

Terry C. Jones

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole H₂S and multiply it with 2 mole H₂O over 2 mole H₂S from the equation

$$8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$$

8.3 mole H₂S x 18g H₂O = 149.4g H₂O is your theoretical yield so percent yield = $\frac{137.1\text{g}}{149.4\text{g}} \times 100 = 91.77\%$

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

if the reaction of 21.3 grams of O₂ produces a 47.3% yield how many grams of CO would be produced

step one figure out how many moles of the limiting reagent you have

$$21.3 \text{ g O}_2 \times \frac{1 \text{ mole}}{32 \text{ g}} = 0.666 \text{ moles O}_2$$

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web percent yield is a measure of the efficiency of the experimental design

$$\text{yield} = \frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$$

in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction

$$2\text{Al} + 3\text{O}_2 \rightarrow \text{Al}_2\text{O}_3$$

b determine the theoretical yield of Al₂O₃

$$160.0 \text{ g Al} \times \frac{1 \text{ mole Al}}{27 \text{ g}} \times \frac{2 \text{ mole Al}_2\text{O}_3}{4 \text{ mole Al}} \times 101.96 \text{ g} = 160.0 \text{ g Al}_2\text{O}_3$$

c determine the percent yield

$$\frac{160.0 \text{ g}}{260.0 \text{ g}} \times 100 = 61.5\%$$

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