

A student investigates the evaporation of water. He pours 100cm^3 of water into measuring cylinder A and 100cm^3 of water into measuring cylinder B. Measuring cylinder A is kept at 40°C and B is kept at 80°C in the same part of the laboratory. Fig. 2.1 shows the two measuring cylinders after 3 days.



Fig. 2.1

- (a) (i) Using ideas about molecules, explain why more water evaporates when the water is kept at 80°C .

.....

 [2]

- (ii) Apart from an increase in temperature, state one change that causes water to evaporate faster.

.....
 [1]

- (b) The specific latent heat of vaporisation of water is 2300J/g and the density of water is 1.0g/cm^3 . During the three days, the water level in B drops from the 100cm^3 mark to the level shown in Fig. 2.1. Calculate the energy used to evaporate water from B during the three days.

energy = [2]

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