

## 7. GROUPS 1 & 2

### SECTION B

Answer ALL the questions. Write your answers in the spaces provided.

- 1 Magnesium oxide, known as magnesia, can be produced by the thermal decomposition of magnesium carbonate.

The equation for the reaction is



- \*(a) Other Group 2 carbonates also undergo thermal decomposition.

State and explain the trend in thermal stability of carbonates down Group 2.

(4)

Thermal stability increases down the group

Because ionic radius increases as you go down the group

This results in less polarisation (distortion) of the carbonate ion

And less weakening of the covalent bonds (C-O) within the carbonate ion

- \*(c) State why carrying out a flame test is not an appropriate way to confirm the presence of magnesium ions in the sample of magnesia.

(1)

Magnesium ions give no flame colour

OR

Energy is emitted with Mg ions is outside of the visible region

(Total for Question 1 = 15 marks)

# Igcse Edexcel Past Papers Biology

**Xiang Xie**



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## **Table of Contents Igcse Edexcel Past Papers Biology**

1. Understanding the eBook Igcse Edexcel Past Papers Biology
  - The Rise of Digital Reading Igcse Edexcel Past Papers Biology
  - Advantages of eBooks Over Traditional Books
2. Identifying Igcse Edexcel Past Papers Biology
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Igcse Edexcel Past Papers Biology
  - User-Friendly Interface
4. Exploring eBook Recommendations from Igcse Edexcel Past Papers Biology
  - Personalized Recommendations
  - Igcse Edexcel Past Papers Biology User Reviews and Ratings
  - Igcse Edexcel Past Papers Biology and Bestseller Lists
5. Accessing Igcse Edexcel Past Papers Biology Free and Paid eBooks
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  - Igcse Edexcel Past Papers Biology eBook Subscription Services
  - Igcse Edexcel Past Papers Biology Budget-Friendly Options
6. Navigating Igcse Edexcel Past Papers Biology eBook Formats
  - ePub, PDF, MOBI, and More
  - Igcse Edexcel Past Papers Biology Compatibility with Devices
  - Igcse Edexcel Past Papers Biology Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Igcse Edexcel Past Papers Biology
  - Highlighting and Note-Taking Igcse Edexcel Past Papers Biology
  - Interactive Elements Igcse Edexcel Past Papers Biology
8. Staying Engaged with Igcse Edexcel Past Papers Biology

- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Igcse Edexcel Past Papers Biology
9. Balancing eBooks and Physical Books Igcse Edexcel Past Papers Biology
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Igcse Edexcel Past Papers Biology
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Igcse Edexcel Past Papers Biology
- Setting Reading Goals Igcse Edexcel Past Papers Biology
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Igcse Edexcel Past Papers Biology
- Fact-Checking eBook Content of Igcse Edexcel Past Papers Biology
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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