

Percent Yield Problems

Theoretical Yield = The *maximum* amount of product that can be formed from given reactants.

--Theoretical yield is the answer from the stoichiometry calculation = maximum possible amount

---Number based upon the "theoretical" mathematical equation – no experiments were conducted

Actual Yield = The amount of product that *actually* forms when the reaction is carried out in the lab.

--Due to a variety of factors, the actual amount will ALWAYS be less than the theoretical yield.

---In real life, we never get as much as we anticipate, this "actual" amount from the lab experiment

Percent Yield = The *ratio* of the actual yield to the theoretical yield expressed as a percentage.

$$\text{Percent Yield} = \frac{\text{Actual}}{\text{Theoretical}} \times 100$$

EXAMPLE) How many grams of hydrofluoric acid (HF) are required to react completely with 23.68 grams of calcium hydroxide in the following reaction? \square



$$23.68 \text{ g Ca(OH)}_2 \times \frac{1 \text{ mol Ca(OH)}_2}{74.10 \text{ g Ca(OH)}_2} \times \frac{2 \text{ mol HF}}{1 \text{ mol Ca(OH)}_2} \times \frac{20.01 \text{ g HF}}{1 \text{ mol HF}} = \boxed{12.79 \text{ g HF}}$$

Based on the stoichiometry of this problem, the **Theoretical Yield** for this problem is 12.79 g HF. But, when this experiment was conducted, an **Actual Yield** of only 10.41 g HF was collected. Using the equation for percent yield, the **Percent Yield** of this experiment was 81.39%.

$$\text{Percent Yield} = \frac{\text{Actual}}{\text{Theoretical}} \times 100 = \frac{10.41 \text{ g HF}}{12.79 \text{ g HF}} \times 100 = \boxed{81.39\%}$$

1) A student adds 200.0g of $\text{C}_7\text{H}_6\text{O}_3$ to an excess of $\text{C}_2\text{H}_5\text{O}_2$, this produces $\text{C}_9\text{H}_8\text{O}_4$ and $\text{C}_2\text{H}_4\text{O}_2$. Calculate the percent yield if 231 g of aspirin ($\text{C}_9\text{H}_8\text{O}_4$) is produced in an experiment.



2) According to the following equation, Calculate the percentage yield if 550.0 g of toluene (C_7H_8) added to an excess of nitric acid (HNO_3) provides 305 g of the p-nitrotoluene ($\text{C}_7\text{H}_7\text{NO}_2$) product in a lab experiment.



3) Aluminum reacts with an aqueous solution containing excess copper (II) sulfate. If 1.85 g Al reacts and the percentage yield of Cu is 56.6%, what mass of Cu is produced?



4) The combustion of methane(CH_4) produces carbon dioxide and water. Assume that 2.0 mol of CH_4 burned in the presence of excess air. What is the percentage yield if in an experiment the reaction produces 87.0 g of CO_2 ?

5) 15.3 grams of Lithium is dropped into a solution containing excess copper II phosphate. When the reaction is completed, 1.25 grams of copper is formed. What is the percent yield?

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Progress in Ecological Stoichiometry Dedmer B. Van de Waal,James J. Elser, Adam C. Martiny,Robert W. Sterner,James B. Cotner,2018 Ecological stoichiometry concerns the way that the elemental composition of organisms shapes their ecology It deals with the balance or imbalance of elemental ratios and how that affects organism growth nutrient cycling and the interactions with the biotic and abiotic worlds The elemental composition of organisms is a set of constraints through which all the Earth s biogeochemical cycles must pass All organisms consume nutrients and acquire compounds from the environment proportional to their needs Organismal elemental needs are determined in turn by the energy required to live and grow the physical and chemical constraints of their environment and their requirements for relatively large polymeric biomolecules such as RNA DNA lipids and proteins as well as for structural needs including stems bones shells etc These materials together constitute most of the biomass of living organisms Although there may be little variability in elemental ratios of many of these biomolecules changing the proportions of different biomolecules can have important effects on organismal elemental composition Consequently the variation in elemental composition both within and across organisms can be tremendous which has important implications for Earth s biogeochemical cycles It has been over a decade since the publication of Sterner and Elser s book *Ecological Stoichiometry* 2002 In the intervening years hundreds of papers on

stoichiometric topics ranging from evolution and regulation of nutrient content in organisms to the role of stoichiometry in populations communities ecosystems and global biogeochemical dynamics have been published Here we present a collection of contributions from the broad scientific community to highlight recent insights in the field of Ecological Stoichiometry

Student Study Guide Fundamentals of Chemistry by David E. Goldberg Margaret G. Kimble,1994 **Fundamentals of Chemistry, Study Guide** James E. Brady,John R. Holum,1988-04-20 This Third Edition of the widely used fundamentals textbook for science majors maintains the conversational writing style that made the previous editions so popular while including up to date treatments of important and current topics Emphasizes descriptive chemistry chemical reactions and properties while maintaining a solid treatment of chemical principles Common chemicals are used whenever possible as examples in both theoretical discussions and in problems and exercises Incorporates many pedagogical aids each chapter begins with a brief table of contents and each section begins with a preview of topics covered Chapters include frequent margin comments figures and photographs *Study and Problem Solving Guide to Accompany Principles of Modern Chemistry, Oxtoby/Nachtrieb* Wade A. Freeman,1986 Study Guide for General Chemistry and College Chemistry, Eighth Editions by Holtzclaw and Robinson Norman E. Griswold, Henry Fuller Holtzclaw,1988 *General Chemistry, Study Guide* James E. Brady,1990-03-01 The Fifth Edition retains the pedagogical strengths that made the previous editions so popular and has been updated reorganized and streamlined Changes include more accessible introductory chapters with greater stress on the logic of the periodic table earlier introduction of redox reactions greater emphasis on the concept of energy a new section on Lewis structures earlier introduction of the ideal gas law and a new development of thermodynamics Each chapter ends with review questions and problems **Resources in Education** ,1995 Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl Paul B. Kelter,1993 **Popular Science** ,1877-06 Popular Science gives our readers the information and tools to improve their technology and their world The core belief that Popular Science and our readers share The future is going to be better and science and technology are the driving forces that will help make it better The Extraordinary Chemistry of Ordinary Things, Study Guide Carl H. Snyder,David F. Dever,1995-01-19 Examines the chemistry of the substances of our everyday world Our daily lives are immersed in chemicals an effective way to teach and learn chemistry is by examining the goods and substances that we use in our daily lives and that affect us and our environment

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