Key

Molecular Geometry

How can molecular shapes be predicted using the VSEPR theory?

Why?

When you draw a Lewis structure for a molecule on paper, you are making a two-dimensional representation of the atoms. In reality however, molecules are not flat—they are three-dimensional. The true shape of a molecule is important because it determines many physical and chemical properties for the substance. In this activity you will learn how to predict molecular shapes.

Model 1 - Lewis Structures

| Lows Structures | H ₂ CO 3 electron domains (3 bonding, 0 nonbonding) | 3-D Molecular Shape |
|-----------------|---|---------------------|
| 2 PaF; | BeF ₂ 2 electron domains (2 bonding, 0 nonbonding) | E Tetrakedral |
| 3 04 | CH ₄ 4 electron domains (4 bonding, 0 nonbonding) | o go Tetrakedral |
| 4. Nets | NH ₃ 4 electron domains (3 bonding, 1 nonbonding) | 880 pyramio |
| • | H ₃ O 4 electron domains (2 bonding, 2 nonbonding) | a Bent |
| Lone pair * ** | CO ₂ 2 electron domains (2 bonding, 0 nonbonding) | 0=0=0 |

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