

I. The Reaction Quotient

1. We can write K only for a reaction at equilibrium
2. How can we describe a reaction at other times? $\text{N}_2 + 3\text{H}_2 \rightleftharpoons 2\text{NH}_3$

$$Q = \frac{[\text{NH}_3]_0^2}{[\text{N}_2]_0[\text{H}_2]_0^3} \quad \text{Describes the reaction when not at equilibrium}$$

3. $Q = \text{reaction quotient}$ has the same form as K , but is at non-equilibrium times.
4. Q can tell us how a reaction will change to get to equilibrium
 - 1) If $Q = K$, we are at equilibrium
 - 2) If $Q > K$, then products $>$ reactants and the reaction will shift left
 - 3) If $Q < K$, then reactants $>$ products and the reaction will shift right

Q	K	Predicted Direction of Reaction
0.55	1.45	To the right (towards products)
2.55	1.45	To the left (towards reactants)
1.45	1.45	No change (at equilibrium)

5. *Example:* For the above reaction, $K = 6 \times 10^{-2}$. Predict change:
 - 1) $[\text{N}_2]_0 = 1 \times 10^{-5} \text{ M}$, $[\text{H}_2]_0 = 0.002 \text{ M}$, $[\text{NH}_3]_0 = 0.001 \text{ M}$
 - 2) $[\text{N}_2]_0 = 1.5 \times 10^{-5} \text{ M}$, $[\text{H}_2]_0 = 0.354 \text{ M}$, $[\text{NH}_3]_0 = 0.0002 \text{ M}$
 - 3) $[\text{N}_2]_0 = 5 \text{ M}$, $[\text{H}_2]_0 = 0.01 \text{ M}$, $[\text{NH}_3]_0 = 0.0001 \text{ M}$

Entropy Pogil Answer Key

Annelies Wilder-Smith

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

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