

Empirical Formula Practice 1.0 – Answer Key

Objective: – Calculate the empirical formula of a compound given experimental data

Directions: Answer the following empirical formula questions. Show all work! Include units!

1. Dibutyl succinate is an insect repellent used against household ants and roaches. Its composition is 62.58% C, 9.63% H and 27.79% O. Find its empirical formula. $[C_{16}H_{22}O_4]$

$$C: = 5.210658 \text{ mol} + 1.736875 = 6.947533 \quad = 3 \quad \times 2 = 6$$

$$H: = 9.534653 \text{ mol} + 1.736875 = 11.271528 \quad = 5.5 \quad \times 2 = 11$$

$$O: = 1.736875 \text{ mol} + 1.736875 = 3.47375 \quad = 1 \quad \times 2 = 2$$



2. The compound adrenaline contains 56.7% C, 6.56% H, 8.28% N, and 28.37% O by mass. What is the empirical formula for adrenaline?

$$C: = 4.721056 \text{ mol} + 0.591006 = 5.312062 \quad = 8$$

$$H: = 6.455050 \text{ mol} + 0.591006 = 7.046056 \quad = 11$$

$$N: = 0.591006 \text{ mol} + 0.591006 = 1.182012 \quad = 1$$

$$O: = 1.773125 \text{ mol} + 0.591006 = 2.364131 \quad = 3$$



3. A sample of urea contains 0.480 g of C, 0.161 g of H, 1.121 g of N, and 0.640 g of O. What is the empirical formula of urea?

$$C: = 0.039967 \text{ mol} + 0.039967 = 0.079934 \quad = 1$$

$$H: = 0.159406 \text{ mol} + 0.039967 = 0.199373 \quad = 4$$

$$N: = 0.080014 \text{ mol} + 0.039967 = 0.119981 \quad = 2$$

$$O: = 0.040000 \text{ mol} + 0.039967 = 0.079967 \quad = 1$$



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Chemistry: The Central Science Theodore L. Brown, H. Eugene LeMay Jr., Bruce E. Bursten, Catherine Murphy, Patrick Woodward, Steven Langford, Dalius Sagatys, Adrian George, 2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet retains the clarity innovative pedagogy functional problem solving and visuals of the previous version All artwork and images are now consistent in quality across the entire text And with a more traditional and logical organisation of the Organic Chemistry content this comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding development of problem solving skills reference and test preparation

Chemistry in a Month John Wilkinson, 2003

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General Chemistry Donald A. McQuarrie, Stanley Gill, 2011-06-15 This Fourth Edition of McQuarrie s classic text offers a thorough revision and a quantum leap forward from the previous edition Taking an atoms first approach it promises to be another ground breaking text in the tradition of

McQuarrie's many previous works. This outstanding new text, available in a soft cover edition, offers professors a fresh choice and outstanding value. **33 Years NEET Chapterwise & Topicwise Solved Papers CHEMISTRY (2020 - 1988) 15th Edition** Disha Experts, *Basics for Chemistry* David A. Ucko, 2013-09-24. Basics of Chemistry provides the tools needed in the study of General Chemistry such as problem solving skills, calculation methods, and the language and basic concepts of chemistry. The book is designed to meet the specific needs of underprepared students. Concepts are presented only as they are needed and developed from the simple to the complex. The text is divided into 18 chapters, each covering some particular aspect of chemistry such as matter, energy, and measurement; the properties of atoms; description of chemical bonding; study of chemical change; and nuclear and organic chemistry. Undergraduate students will find the book a very valuable academic material.

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Table of Contents Empirical Formula Key Answers

1. Understanding the eBook Empirical Formula Key Answers
 - The Rise of Digital Reading Empirical Formula Key Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Empirical Formula Key Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Empirical Formula Key Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Empirical Formula Key Answers
 - Personalized Recommendations
 - Empirical Formula Key Answers User Reviews and Ratings
 - Empirical Formula Key Answers and Bestseller Lists
5. Accessing Empirical Formula Key Answers Free and Paid eBooks
 - Empirical Formula Key Answers Public Domain eBooks
 - Empirical Formula Key Answers eBook Subscription Services
 - Empirical Formula Key Answers Budget-Friendly Options
6. Navigating Empirical Formula Key Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Empirical Formula Key Answers Compatibility with Devices
 - Empirical Formula Key Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Empirical Formula Key Answers
 - Highlighting and Note-Taking Empirical Formula Key Answers
 - Interactive Elements Empirical Formula Key Answers
8. Staying Engaged with Empirical Formula Key Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Empirical Formula Key Answers
- 9. Balancing eBooks and Physical Books Empirical Formula Key Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Empirical Formula Key Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Empirical Formula Key Answers
 - Setting Reading Goals Empirical Formula Key Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Empirical Formula Key Answers
 - Fact-Checking eBook Content of Empirical Formula Key Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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