

2. Magnesium carbonate reacts with hydrochloric acid.



- (a) What mass of magnesium carbonate would react with excess hydrochloric acid to produce 240cm^3 of carbon dioxide, measured at room temperature and pressure?

Data: 1 mol of any gas occupies 24.0dm^3 at room temperature and pressure

Molar mass of magnesium carbonate = 84.3g mol^{-1}

- ☒ A 0.843 g
☐ B 8.43 g
☐ C 84.3 g
☐ D 843 g

$$\begin{aligned}\text{Volume of CO}_2 &= 240\text{cm}^3 = 0.24\text{dm}^3 \\ \text{moles of CO}_2 &= \frac{\text{Volume in dm}^3}{24} = \frac{0.24}{24} = 0.01 \\ \text{MgCO}_3 &\text{ contains same moles as CO}_2 \\ \text{moles} &= \frac{\text{mass}}{\text{Mr}} ; \text{ mass} = \text{moles} \times \text{Mr} \\ &= 0.01 \times 84.3 \\ &= 0.843\text{g}\end{aligned}$$

- (b) What is the minimum mass of magnesium carbonate needed to neutralise 50.0cm^3 of 0.250mol dm^{-3} hydrochloric acid?

- ☐ A 0.423 g
☒ B 0.527 g
☐ C 1.05 g
☐ D 2.11 g

$$\begin{aligned}\text{Concentration of HCl} &= \frac{\text{moles}}{\text{Vol in dm}^3} \\ \text{moles} &= \text{Conc} \times \text{Vol in dm}^3 \\ &= 0.250 \times 0.05 \\ &= 0.0125\text{moles of HCl}\end{aligned}$$

$$\begin{aligned}\text{MgCO}_3 : \text{HCl} \\ 1 : 2 \\ \frac{1}{2} : 1 \\ \frac{1}{2} \times 0.0125 : 0.0125 \\ 6.25 \times 10^{-3} : 0.0125 \\ \downarrow \qquad \qquad \downarrow \\ \text{MgCO}_3 \qquad \text{HCl} \\ \text{moles} \qquad \text{moles}\end{aligned}$$

- (c) What would be seen at the end of the reaction with excess acid?

- ☒ A A colourless solution
☐ B A coloured solution
☐ C A white precipitate
☐ D A coloured precipitate

all MgCO_3 is utilized
and color of
 MgCO_2 colorless

$$\text{mass} = \text{moles} \times \text{Mr}$$

(Total for Question 2 = 3 marks)

$$\begin{aligned}&= 6.25 \times 10^{-3} \times 84.3 \\ &= 0.527\text{g}\end{aligned}$$

Multiple Choice Questions On Stoichiometry

M Mosston



Multiple Choice Questions On Stoichiometry:

This book delves into Multiple Choice Questions On Stoichiometry. Multiple Choice Questions On Stoichiometry is an essential topic that must be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Multiple Choice Questions On Stoichiometry, encompassing both the fundamentals and more intricate discussions.

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