

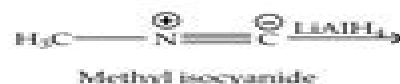
SOLUTIONS

Specimen Paper 2023

Chemistry

Section-A

1. (A) (i) anisole, Williamson's synthesis.
 (ii) three, s^{-1}
 (iii) lewis acid, lewis base.
 (iv) acetone, iodoform.
 (B) (i) Option (b) is correct.
Explanation:



- (ii) Option (c) is correct.

Explanation: For a spontaneous reaction, the standard change in standard Gibbs free energy, ΔG° and the standard cell potential, E_{cell}° will be negative and positive respectively.

- (iii) Option (c) is correct.

Explanation: Electronic configuration of $\text{Cr} = [\text{Ar}] 3d^5 4s^1$ and $\text{Cu} = [\text{Ar}] 3d^{10} 4s^1$

- (iv) Option (a) is correct.

Explanation: Since, for first order reaction, rate law is expressed as

$$k = \frac{2.303}{t} \log \frac{[R_0]}{[R]}$$

At $t = 8 \text{ min}$,

$$k = \frac{2.303}{8} \log \frac{100}{25} = 0.1733 \text{ min}^{-1}$$

For 200 g,

$$k = \frac{2.303}{0.1733} \log \frac{200}{50} = 8 \text{ min}$$

- (C) (i) Phenol
 (c) Neutral FeCl_3 solution
 (ii) Ethylenediamine
 (d) Bidentate ligand
 (iii) Colligative property
 (a) Osmotic pressure
 (iv) Amino acid
 (b) Zwitter ion

- (D) (i) Option (a) is correct.

Explanation: Conductivity always decreases with the decrease in concentration for both weak and strong electrolytes as the number of ions per unit volume that carry the current

in a solution decreases on dilution.

- (ii) Option (d) is correct.

Explanation: Chlorination of nitrobenzene (not the nitration of chlorobenzene) leads to the formation of *m*-nitrochlorobenzene because $-\text{NO}_2$ group deactivates the ring because it is meta directing.

Section-B

2. Osmotic pressure is a colligative property. According to it,

$$\pi = CRT$$

$$\pi = \frac{n}{V} RT \quad \left(C = \frac{n}{V} = \frac{m}{MV} \right)$$

m = Mass of solute = 20 g

M = Molar mass

V = volume of solution = 500 mL

= 0.5 L

R = Gas constant = 0.0821 atm

$\text{mol}^{-1} \text{K}^{-1}$

T = Temperature = 298 K

Place the values in the above equation

$$\pi = \frac{m}{MV} RT$$

$$0.016 = \frac{20 \times 1000}{M \times 500} \times 0.0821 \times 298$$

$$M = \frac{20 \times 1000}{0.016 \times 500} \times 0.0821 \times 298$$

$$= 61164.5 \text{ g}$$

3. (i) Transition elements form a large number of complex compounds due to the comparatively smaller sizes of the metal ions and their high ionic charges and also the availability of

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