KEY

FINAL EXAM PRACTICE SHEET

- Write the molecular, total ionic, and net ionic equations of the following reactions
 (15pts): + predict the products
- a. Zinc metal is placed in aqueous hydrochloric acid

ME
$$Z_{n}(s) + 2H(e_{qq}) \rightarrow Z_{n}(e_{q}) + H_{2}(q)$$

TE $Z_{n}(s) + 2H'e_{q}) + 2CC'e_{q}) \rightarrow Z_{n}^{2+}(e_{q}) + 2CC'(e_{q}) + H_{2}(q)$
NIE $Z_{n}(s) + 2H'e_{q}) \rightarrow Z_{n}^{2+}(e_{q}) + H_{2}(q)$

b. Aqueous barium chloride is mixed with aqueous sodium sulfate

ME Bally + Nay SO4 (an)
$$\rightarrow$$
 BaSO4 (as) + 2 Nall (an)

IE Baten) + 20 (an) + 2 Nat (an) + 50 to (an) \rightarrow BaSO4 (s) + 2 Nat (an)

+ 20 (an)

NIE Baten) + 50 to (an) \rightarrow BaSO4 (s)

c. Aqueous ammonium sulfate is mixed with aqueous potassium hydroxide

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Lingsheng Yao

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Table of Contents Intro To Chem Final Exam

- 1. Understanding the eBook Intro To Chem Final Exam
 - The Rise of Digital Reading Intro To Chem Final Exam
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Intro To Chem Final Exam
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Intro To Chem Final Exam
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Intro To Chem Final Exam
 - Personalized Recommendations
 - Intro To Chem Final Exam User Reviews and Ratings
 - Intro To Chem Final Exam and Bestseller Lists
- 5. Accessing Intro To Chem Final Exam Free and Paid eBooks
 - o Intro To Chem Final Exam Public Domain eBooks
 - Intro To Chem Final Exam eBook Subscription Services
 - Intro To Chem Final Exam Budget-Friendly Options

- 6. Navigating Intro To Chem Final Exam eBook Formats
 - o ePub, PDF, MOBI, and More
 - Intro To Chem Final Exam Compatibility with Devices
 - Intro To Chem Final Exam Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Intro To Chem Final Exam
 - Highlighting and Note-Taking Intro To Chem Final Exam
 - Interactive Elements Intro To Chem Final Exam
- 8. Staying Engaged with Intro To Chem Final Exam
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Intro To Chem Final Exam
- 9. Balancing eBooks and Physical Books Intro To Chem Final Exam
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Intro To Chem Final Exam
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Intro To Chem Final Exam
 - Setting Reading Goals Intro To Chem Final Exam
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Intro To Chem Final Exam
 - Fact-Checking eBook Content of Intro To Chem Final Exam
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements

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