

# Intro to Acids and Bases

Name: \_\_\_\_\_  
 Class: \_\_\_\_\_  
 Date: \_\_\_\_\_

## Define each term: Definitions of acids and bases:

### Arrhenius Definitions:

- acid substances that release dissolved hydrogen ions ( $H^+$ ) (acids  $H^+$  gases formed in gases as the hydrogens form  $H_2$  and  $H_2O$  (acid  $H_2O$ ) and conjugated anions)
- base substances which dissociate water molecules ( $H_2O$ )

### Brønsted-Lowry Definitions:

- acid substances that donate a proton ( $H^+$ ) (acid donate)
- base substances which accept a proton ( $H^+$ ) (base receive)

Table 1: Representative reactions for  $H^+$  (acid) water with base representative reactions:

Reaction	Eq.
1. $CH_3COOH + H_2O \rightleftharpoons CH_3COO^- + H_3O^+$	$CH_3COOH + H_2O$
2. $HNO_3 + H_2O \rightleftharpoons HNO_2 + H_3O^+$ (acidic) $\rightarrow H_2O + H_2O$	$HNO_3 + H_2O$
3. $HNO_3 + H_2O \rightleftharpoons HNO_2 + H_3O^+$ (acidic) $\rightarrow H_2O + H_2O$	$HNO_3 + H_2O$
4. $HNO_3 + H_2O \rightleftharpoons HNO_2 + H_3O^+$ (acidic) $\rightarrow H_2O + H_2O$	$HNO_3 + H_2O$
5. $HNO_3 + H_2O \rightleftharpoons HNO_2 + H_3O^+$ (acidic) $\rightarrow H_2O + H_2O$	$HNO_3 + H_2O$

## Critical Thinking Questions

- Are all the substances in Table 1 that are representative acids?
- Are all the substances in Table 1 that are representative bases?

### Class work:

- Are all the substances in Table 1 that are Brønsted-Lowry acids?

### Individual questions for students:

- Are all the substances in Table 1 that are Brønsted-Lowry bases?

### Group work questions:

- Are all the substances in Table 1 that are Brønsted-Lowry acids?

### Individual questions for students:

- Are all the substances in Table 1 that are Brønsted-Lowry bases?

### Class:

- Are all the substances in Table 1 that are Brønsted-Lowry acids?

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