

2. Magnesium carbonate reacts with hydrochloric acid.



- (a) What mass of magnesium carbonate would react with excess hydrochloric acid to produce  $240\text{cm}^3$  of carbon dioxide, measured at room temperature and pressure?

Data: 1 mol of any gas occupies  $24.0\text{dm}^3$  at room temperature and pressure

Molar mass of magnesium carbonate =  $84.3\text{g mol}^{-1}$

- ☒ A 0.843 g  
☐ B 8.43 g  
☐ C 84.3 g  
☐ D 843 g

$$\begin{aligned} \text{Volume of CO}_2 &= 240\text{cm}^3 = 0.24\text{dm}^3 \\ \text{moles of CO}_2 &= \frac{\text{Volume in dm}^3}{24} = \frac{0.24}{24} = 0.01 \\ \text{MgCO}_3 &\text{ contains same moles as CO}_2 \\ \text{moles} &= \frac{\text{mass}}{\text{Mr}} ; \text{ mass} = \text{moles} \times \text{Mr} \\ &= 0.01 \times 84.3 \\ &= 0.843\text{g} \end{aligned}$$

- (b) What is the minimum mass of magnesium carbonate needed to neutralise  $50.0\text{cm}^3$  of  $0.250\text{mol dm}^{-3}$  hydrochloric acid?

- ☐ A 0.423 g  
☒ B 0.527 g  
☐ C 1.05 g  
☐ D 2.11 g

$$\begin{aligned} \text{Concentration of HCl} &= \frac{\text{moles}}{\text{Vol in dm}^3} \\ \text{moles} &= \text{Conc} \times \text{Vol in dm}^3 \\ &= 0.250 \times 0.05 \\ &= 0.0125\text{moles of HCl} \end{aligned}$$

$$\begin{aligned} \text{MgCO}_3 : \text{HCl} \\ 1 : 2 \\ \frac{1}{2} : 1 \\ \frac{1}{2} \times 0.0125 : 0.0125 \\ 6.25 \times 10^{-3} : 0.0125 \\ \downarrow \qquad \qquad \downarrow \\ \text{MgCO}_3 \qquad \text{HCl} \\ \text{moles} \qquad \text{moles} \end{aligned}$$

- (c) What would be seen at the end of the reaction with excess acid?

- ☒ A A colourless solution  
☐ B A coloured solution  
☐ C A white precipitate  
☐ D A coloured precipitate

all  $\text{MgCO}_3$  is utilized  
and color of  
 $\text{MgCO}_2$  colorless

$$\text{mass} = \text{moles} \times \text{Mr}$$

(Total for Question 2 = 3 marks)

$$\begin{aligned} &= 6.25 \times 10^{-3} \times 84.3 \\ &= 0.527\text{g} \end{aligned}$$

# Multiple Choice Questions Of Stoichiometry

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## Multiple Choice Questions Of Stoichiometry:

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Chapadgaonkar, Anamika Singh, 2019-10-17 Biochemistry is the study of chemical processes within and relating to living organisms Biochemical processes give rise to complexity of life Biochemistry deals with the structures functions and interactions of biological macromolecules such as proteins nucleic acids carbohydrates and lipids which provide the structure of cells and perform many functions of life This book offers a comprehensive coverage of biochemical systems structures and metabolic reactions within the organism while emphasizing on the nutritional aspects of biochemicals The chapters are designed such that it covers all the concepts from the basic to the complex to help the students understand key concepts and appreciate the significant connections among biochemistry physiology nutrition and human health In this text individual disciplines are treated in an integrated manner to provide sound foundation in the subject The book is particularly meant for the students from undergraduate and postgraduate courses in life sciences health sciences pharmaceutical sciences and other related fields since it gives fundamental knowledge of biochemistry and explains well the metabolic and physiological roles of biomolecules

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