

ENERGY CHANGES IN PHYSICAL AND CHEMICAL PROCESSES.

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Objectives

By the end of this topic, the learner should be able to:

- Define endothermic and exothermic reactions using the ΔH notation.
- Draw energy level and energy cycle diagrams.
- Explain fusion and vaporisation as evidence of inter-particle forces.
- Explain that energy changes in chemical reactions are due to bond breaking and bond formation.
- Define and explain various types of heat changes.
- Carry out some experiments to determine enthalpy changes for some reactions.
- Write correct simple thermo-chemical equations.
- State Hess's Law and carry out related calculations.
- State and explain the factors that influence the choice of a fuel.
- Explain the environmental effects of fuels.

ENERGY CHANGES IN PHYSICAL AND CHEMICAL PROCESSES.

Energy is the ability to do work. There are many forms of energy such as chemical, electrical, heat, kinetic and potential energy.

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Willy Kirkeby



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