16

SOLUTIONS

SECTION 16.1 PROPERTIES OF SOLUTIONS (pages 471–477)

This section identifies the factors that affect the solubility of a substance and determine the sute at which a solute dissolves.

► Solution Formation (pages 471-472)

Look at Figure 16.1 on page 471 to help you answer Questions 1 and 2.

1. Underline the condition that causes sugar to dissolve faster in water.

a. as a whole cube

or in granulated form?

b. when allowed to stand

or when stirred?

c. at a higher temperature

or a lower temperature?

Name three factors that influence the rate at which a solute dissolves in a solvent.

agitation (stirring or shaking)

h. temperature

particle size

3. Is the following sentence true or false? Finely ground particles dissolve more rapidly than larger particles because finer particles expose a greater surface area to the colliding solvent molecules.

► Solubility (pages 472-473)

 Complete the following table showing the steps in a procedure to determine the total amount of sodium chloride that will dissolve in 100 g of water at 25°C.

Procedure	Amount Dissolved	Amount Not Dissolved
Add 36.0 g of sodium chloride to the water	36.0 g	0.0 g
Add an additional 1.0 g of sodium chloride	0.2 g	0.8 g
Determine the total amount that dissolves	36.2 g	

- The amount of a substance that dissolves in a given quantity of solvent at a
 constant temperature is called the substance's ______ at that temperature.
- If a solution contains the maximum amount of solute for a given quantity of solvent at a constant temperature, it is called a(n) <u>saturated</u> solution.

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