



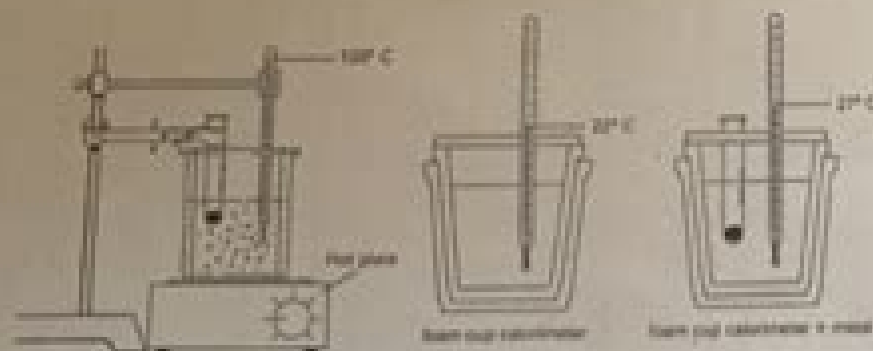
INTERPRETING GRAPHICS

USE WITH SECTION 11.1

A student performed an experiment to determine the specific heat of an unknown metal. The data she collected is reproduced in the table below. Use this information to answer the following questions.

Quantity	Trial 1	Trial 2
1. Mass of test tube + metal	118.19 g	118.21 g
2. Mass of test tube	47.88 g	47.88 g
3. Mass of metal	70.31 g	70.33 g
4. Mass of calorimeter	1.237 g	1.218 g
5. Mass of calorimeter + water	46.127 g	45.719 g
6. Mass of water	38.100 g	38.452 g
7. Initial temperature of metal	100.0 °C	100.0 °C
8. Initial temperature of water	22.8 °C	21.6 °C
9. Final temperature of water	27.8 °C	26.1 °C

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1. What was the final temperature of the metal?

a. Trial 1

b. Trial 2

2. What was the ΔT for the metal?

a. Trial 1

b. Trial 2

3. What was the ΔT for the water?

a. Trial 1

b. Trial 2

4. Calculate the heat change for the water. (Specific heat capacity for water is $4.184 \text{ J/g}^\circ\text{C}$)

a. Trial 1

b. Trial 2

5. Calculate the heat change for the metal.

a. Trial 1

b. Trial 2

6. Calculate the specific heat of the metal.

a. Trial 1

b. Trial 2

7. What metal might this have been?

a. aluminum

b. silver

c. iron

d. mercury

88 Review Module 1 (Chapters 8-12)

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M Woodhall



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