

5

## INTERPRETING GRAPHICS

Use with Section 5.3



Figure 1 The emission spectrum and orbit-transition diagram for hydrogen.

Table 1

Transition	$E(J)$	$r(m)$	$\lambda(m)$	Type of Radiation
$n=6 \rightarrow n=5$	$2.66 \times 10^{-20}$	$4.01 \times 10^{10}$	$7.48 \times 10^{-6}$	Infrared
$n=6 \rightarrow n=4$	$7.57 \times 10^{-21}$	$1.14 \times 10^{10}$	$2.63 \times 10^{-5}$	Infrared
$n=6 \rightarrow n=3$	$1.82 \times 10^{-20}$	$2.15 \times 10^{10}$	$1.09 \times 10^{-5}$	Infrared
$n=6 \rightarrow n=2$	$4.84 \times 10^{-21}$	$7.30 \times 10^{10}$	$4.11 \times 10^{-7}$	Infrared
$n=6 \rightarrow n=1$	$2.12 \times 10^{-20}$	$3.10 \times 10^{10}$	$9.37 \times 10^{-8}$	Visible
$n=5 \rightarrow n=4$	$4.91 \times 10^{-21}$	$7.46 \times 10^{10}$	$4.01 \times 10^{-6}$	Infrared
$n=5 \rightarrow n=3$	$1.25 \times 10^{-20}$	$2.34 \times 10^{10}$	$1.28 \times 10^{-6}$	Infrared
$n=5 \rightarrow n=2$	$4.56 \times 10^{-21}$	$6.88 \times 10^{10}$	$4.36 \times 10^{-7}$	Infrared
$n=5 \rightarrow n=1$	$2.09 \times 10^{-20}$	$3.15 \times 10^{10}$	$9.52 \times 10^{-8}$	Visible
$n=4 \rightarrow n=3$	$1.06 \times 10^{-20}$	$1.60 \times 10^{10}$	$1.87 \times 10^{-6}$	Infrared
$n=4 \rightarrow n=2$	$4.09 \times 10^{-21}$	$6.17 \times 10^{10}$	$4.86 \times 10^{-7}$	Infrared
$n=4 \rightarrow n=1$	$2.04 \times 10^{-20}$	$3.08 \times 10^{10}$	$9.72 \times 10^{-8}$	Visible
$n=3 \rightarrow n=2$	$3.03 \times 10^{-21}$	$4.58 \times 10^{10}$	$6.53 \times 10^{-7}$	Infrared
$n=3 \rightarrow n=1$	$1.94 \times 10^{-20}$	$2.43 \times 10^{10}$	$1.02 \times 10^{-7}$	Visible
$n=2 \rightarrow n=1$	$1.64 \times 10^{-20}$	$2.48 \times 10^{10}$	$1.20 \times 10^{-7}$	Visible

$$\frac{2.18 \times 10^{-18} \text{ J}}{6.626 \times 10^{-34} \text{ J}\cdot\text{s}} = \nu$$

$$\nu = 3.29 \times 10^{15} \text{ s}^{-1}$$

$$h = 6.626 \times 10^{-34} \text{ J}\cdot\text{s}$$

$$c = 2.998 \times 10^8 \text{ m/s}$$

$$c = \nu \lambda$$

$$E = h\nu$$

$$2.998 \times 10^8 =$$

# Interpreting Graphics Ch 10 Chemistry Answers

**CH Cherryholmes**



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