

Cheat Sheet

- molar mass is for a compound (**g/mol**), e.g. $\text{CO}_2 = 44.01 \text{ g/mol}$
- atomic mass is for an element (**g/mol**), e.g. $\text{Al} = 26.98 \text{ g/mol}$

Scenario 1: calculating moles from grams

$$\text{grams of MX} \times \frac{1 \text{ mol of MX}}{\text{grams of MX}} \quad (\text{divide by the molar mass}); \text{ the molar mass isn't affected by the amount of moles in the chemical equation}$$

$$\text{grams of X} \times \frac{1 \text{ mol of X}}{\text{grams of X}} \quad (\text{divide by the atomic mass}); \text{ the atomic mass isn't affected by the amount of moles in the chemical equation}$$

$$88.02 \text{ g of CO}_2 \times \frac{1 \text{ mol CO}_2}{44.01 \text{ g of CO}_2} = 2 \text{ moles of CO}_2$$

Scenario 2: calculating grams from moles

$$\text{moles of MX} \times \frac{\text{grams of MX}}{1 \text{ mol of MX}} \quad (\text{multiply by the molar mass}); \text{ the molar mass isn't affected by the amount of moles in the chemical equation}$$

$$\text{moles of X} \times \frac{\text{grams of X}}{1 \text{ mol of X}} \quad (\text{multiply by the atomic mass}); \text{ the atomic mass isn't affected by the amount of moles in the chemical equation}$$

$$196.97 \text{ moles of Au} \times \frac{196.97 \text{ g of Au}}{1 \text{ mol of Au}} = 38797 \text{ g of Au}$$

Learning Guide To Stoichiometry

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