

Iodine Clock Reaction – Notes for Instructors

Ministry Expectations and Application to Energy Changes and Rates of Reaction Unit

This lab is designed to address a student's understanding of rates of reaction as outlined in the curriculum documents. More specifically, it falls under section "EO: demonstrate an understanding of energy changes and rates of reaction". This experiment involves students mixing standard solutions and monitoring how the rate changes over time with consecutive dilutions. Specifically, this lab demonstrates how concentration can affect the rate of a reaction. Students use the data they collect during the lab to determine the order of the reaction based on their results.

The iodine clock reaction involves mixing two colorless solutions together and then monitoring the solution for a color change. The equations occurring in this experiment are:



In this process, iodide ion is generated by the slow reaction between the iodate and bisulfite. Since this is the slowest step of the reaction, it is also the rate-determining step. The iodate ion excess will oxidize the iodide generated forming iodine. The iodine is immediately then reacts with the starch to form a blue black colored solution. The reaction is sensitive to dissolved oxygen and hence reactions using stock solutions should be carried out fairly soon after they have been prepared. If the distilled water that is used is de-aerated, the solutions will last longer.

Safety Considerations for the Instructor

The preparation for this lab can be done on a bench top. None of the materials are volatile or produce fumes. H₂SO₄ is a strong acid and gloves and goggles should be worn when handling this chemical. Since only a very small amount is required in this lab. Should any contact occur with the skin or eyes rinse with copious amounts of water and consult a doctor as soon as possible.

Safety Considerations for the Students

Students should wear goggles and protective clothing such as a lab coat when performing this experiment. Ensure that no students drink any of the solutions. Should a student come in contact with any of the solutions prepared for this experiment, either on their skin or in their eyes, rinse the affected area with copious amounts of water and go to a hospital immediately.

Preparing the Solutions

NOTE: Water should be de-aerated for H₂SO₄. Hence it is best to pour about 200 or 300 ml into a large beaker so that the 4 liter bottle is tightly capped. This will avoid dissolving oxygen into the stock water.

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