

EXPERIMENT 7
CHEMICAL KINETICS
SAMPLE REPORT

Special Notes

1. The report must be written in ink (blue pen), no pencil allowed. However, the graph can be drawn in pencil.
2. The layout of the report must be tidy (no white out) and legible in writing.
(Work out all the solutions on some scrap papers first, then copy to the report sheets.)
3. Page 7 - 1
 - ① Record exact temperatures in °C to 1 decimal.
 - ② Both the "Time" in seconds and "Relative Rate" must be integers.
 - ③ Report the "Initial [I⁻]", "Initial [S₂O₈²⁻]" and "Rate" in *3 s.f.*
4. Page 7 - 2
 - ① Show **all** calculations for:
 - (A) Concentration of initial [I⁻],
 - (B) Concentration of initial [S₂O₈²⁻],
 - (C) Rate of Reaction of S₂O₈²⁻, and
 - (D) Relative Rates
 - ② Keep all calculations in *3 s.f.*. However, Relative Rate, Order of Reaction, and "m" and "n" must be reported as integers.
5. Page 7 - 3
 - ① For Question 1, the conclusion must include: **rate** = [I⁻].
 - ② For Question 2, the conclusion must include: **rate** = [S₂O₈²⁻].
 - ③ For Question 3, the rate law is: **rate** = **k** [I⁻] [S₂O₈²⁻]
6. Page 7 - 4
 - ① Show calculations of rate constants, **k**, for all five reactions.
 - ② The **units** of rate constant, **k**, is M⁻¹·s⁻¹.
 - ③ The **accuracy** of **k** is between 3.30 × 10⁻³ to 3.90 × 10⁻³ M⁻¹·s⁻¹ at 22-23°C.
 - ④ The **precision** of **k** is ±0.30 × 10⁻³ M⁻¹·s⁻¹ between highest and lowest **k**.

Kinetics Of A Chemical Reaction Lab Report

Padhraic Smyth



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