

## Problem set 2

1. Select the correct alternative and fill in the blank in the following statements.
- (i) If a transversal intersects two parallel lines, the sum of interior angles on the same side of the transversal is .....
- (A)  $0^\circ$  (B)  $90^\circ$  (C)  $180^\circ$  (D)  $360^\circ$
- (ii) The number of angles formed by a transversal on two lines is .....
- (A) 2 (B) 4 (C) 8 (D) 16
- (iii) A transversal intersects two parallel lines. If the measure of one of the angles is  $40^\circ$  then the measure of its corresponding angle is .....
- (A)  $40^\circ$  (B)  $140^\circ$  (C)  $50^\circ$  (D)  $150^\circ$
- (iv) In  $\Delta ABC$ ,  $\angle A = 76^\circ$ ,  $\angle B = 48^\circ$ ,  $\therefore \angle C = \dots$
- (A)  $66^\circ$  (B)  $56^\circ$  (C)  $124^\circ$  (D)  $28^\circ$
- (v) Two parallel lines are intersected by a transversal. If the measure of one of the alternate interior angles is  $75^\circ$  then the measure of the other alternate interior angle is .....

# Geometry Problem Set 2 Solutions 9th Std

**Yan Bai**



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web during oxidation and reduction the oxidation numbers of atoms change if an atom is oxidized its oxidation number increases ie it becomes more ve or less ve if an atom is reduced its oxidation number decreases ie it becomes less ve or more ve

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from one species to another the species that loses electrons is said to be oxidized while the species that gains electrons is said to be reduced

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you 23  $\text{Fe} + \text{Co} \rightarrow \text{Fe}^{2+} + \text{Co}^{2+}$  oxidation  $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^{-}$  reduction  $\text{Fe}^{2+} + 2\text{e}^{-} \rightarrow \text{Fe}$  24 3  $\text{Ag} + \text{Ni} \rightarrow \text{Ag} + \text{Ni}^{3+}$  3  $\text{Ag} \rightarrow \text{Ag}^{+} + \text{e}^{-}$  oxidation reduction 25  $\text{Cu} + \text{Pb} \rightarrow \text{Cu}^{2+} + \text{Pb}$  oxidation reduction 26  $\text{O}_2 \rightarrow \text{O}^{2-}$

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