

MATHEMATICS N2 CLASS EXERCISE 1

QUESTION 1

- 1.1 Solve the following equations:
- 1.1.1 $-x^2 + 3x - 2 = 0$ by factorisation (3)
- 1.1.2 $6(2 - m) - 5(4 - m) = 7$ (3)
- 1.2 Given: the formula $p = \sqrt[3]{r^2 - q^2}$
- 1.2.1 Make ' r ' the subject of the formula (3)
- 1.2.2 Calculate the value of ' r ' if $p = 4$ and $q = 6$ (1)
- 1.3 Solve for ' x ' and ' y ' in the following simultaneous equations:
- $5x = 6y + 3$ and $3y - 4x + 1 = 0$ (4)
- [14]

QUESTION 2

- 2.1 A man deducted 3 from 5 times his age and got his grandmother's age.
If his grandmother is 97 years old, how old is the man? (2)
- 2.2 Given: $4x^2 - 15x - 47 = 0$
Solve for ' x ' the equation using the quadratic formula. (3)
- 2.3 Given: $H = \frac{7,5U}{273 + V}$
- 2.3.1 Make ' V ' the subject of the formula. (3)
- 2.3.2 Find the value of ' V ' if $U = 0,5$ and $H = 0,005$ (1)
- 2.4 Given: $8p - 1 = -2q$ and $2p + 3q = 4$
Solve for ' p ' and ' q ' (4)
- [13]

TOTAL: 27 MARKS

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for the nitration 4 01 ml of nitric acid 4 00 ml of sulfuric acid and 4 56 ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only 4 56 ml of bromobenzene was used the amount of product predicted to be recovered was 8 7533 g

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in this experiment we will carry out the nitration of bromobenzene bromine substituted on an aromatic ring is an ortho para directing group it will direct an incoming electrophile to the ortho and para positions it is also deactivating meaning that its reactions will be slower than the identical reactions of benzene

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in the nitration of bromobenzene a mixture of sulfuric acid and nitric acid react with one another nitric acid the weaker of the two acids will serve as base to form nitronium ion bromobenzene will attack the nitronium ion by nucleophilic attack

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the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area theory and

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1-bromo-4-nitrobenzene

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study with quizlet and memorize flashcards containing terms like reacting sulfuric and nitric acid together produces the nitronium ion alone nitric acid is not a strong enough electrophile to react with bromobenzene sulfuric acid is added to the nitric acid the reverse rxn can overheat and boil easily includes a claisen adaptor and

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over 5 minutes slowly pipet this bromobenzene into the icy cold nitration mixture HNO_3 3 M H_2SO_4 in the 125 ml flask and stir continue to stir the reaction mixture in the ice bath for 30 minutes do not heat in a new 250 ml erlenmeyer

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on the right the arenium cation has a nitro set with incredibly small 0.6 angstrom bond distances this will result in an optimization failure as the atoms are even close for the programme to accomodate their electrons demonstration of creating the p arenium cation by the nitration of bromobenzene

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the purpose of this lab was to observe where substituents are added on a benzene ring by way of electrophilic aromatic substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion NO_2^+ which acts as the electrophilic species

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nitration of bromobenzene is an electrophilic aromatic substitution reaction the electrophile will be formed from 4.0 ml of nitric acid and 4.0 ml of sulfuric acid and will then react with 4.5 ml of bromobenzene this will produce o-bromonitrobenzene

and p bromonitrobenzene

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