

**Exercice 16**  
**Autonomous Work (2)**  
**(20min)**

13. A 1.00 mol sample of  $\text{NH}_3(\text{aq})$  is added to water. The final volume of the solution is 1.0 L.
- a. Calculate the pH of the solution.

$$\begin{aligned} \text{NH}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) &\rightleftharpoons \text{NH}_4^+(\text{aq}) + \text{OH}^-(\text{aq}) \\ K_b &= 1.8 \times 10^{-5} \\ \text{pH} &= 11.25 \end{aligned}$$

- b. In the solution buffer is there 1) anything?

Yes, solution is buffer, as the pH is close to  $\text{p}K_b$ .

14. What is the pH of a 0.10 M solution of  $\text{NH}_4\text{Cl}$ ? ( $K_b$  of  $\text{NH}_3$  is  $1.8 \times 10^{-5}$ .)

$$\begin{array}{l} \text{NH}_4^+(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{NH}_3(\text{aq}) + \text{H}_3\text{O}^+(\text{aq}) \\ K_a = \frac{[\text{NH}_3][\text{H}_3\text{O}^+]}{[\text{NH}_4^+]} = \frac{K_w}{K_b} = \frac{1.0 \times 10^{-14}}{1.8 \times 10^{-5}} = 5.6 \times 10^{-10} \end{array}$$

	$\text{NH}_4^+(\text{aq})$	$\text{H}_2\text{O}(\text{l})$	$\text{NH}_3(\text{aq})$	$\text{H}_3\text{O}^+(\text{aq})$
IC	0.10		0	0
EC	-x		+x	+x
EC	0.10 - x		x	x

$$K_a = \frac{[\text{NH}_3][\text{H}_3\text{O}^+]}{[\text{NH}_4^+]} = \frac{x \cdot x}{0.10 - x} = 5.6 \times 10^{-10}$$

$$\text{pH} = 5.22$$

15. What is the equilibrium concentration of a solution made up of pH of 9.0 at  $\text{p}K_b$ ?

$$\begin{aligned} \text{pH} &= 9.0 \Rightarrow \text{pOH} = 5.0 \\ \text{pOH} &= 5.0 \Rightarrow [\text{OH}^-] = 10^{-5} \text{ M} \\ \text{pOH} &= 5.0 \Rightarrow [\text{OH}^-] = 10^{-5} \text{ M} \end{aligned}$$

$$K_b = \frac{[\text{NH}_3][\text{OH}^-]}{[\text{NH}_4^+]} = \frac{[\text{NH}_3] \cdot 10^{-5}}{[\text{NH}_4^+]} = 1.8 \times 10^{-5}$$

# Lecture 30 Acids And Bases I

**Emil Abderhalden**



## Lecture 30 Acids And Bases I:

**Text-book of Physiological Chemistry in Thirty Lectures** Emil Abderhalden, 1908      **Lectures on Agricultural Chemistry and Geology** James Finlay Weir Johnston, 1847      **Journal of Chemical Education**, 1926 Includes Report of New England Association of Chemistry Teachers and Proceedings of the Pacific Southwest Association of Chemistry Teachers

**Lectures on Mineralogy** Thomas Egleston, 1872      **Lectures on Explosives** Willoughby Walke, 1891      **Introduction to General Chemistry. A Graded Course of One Hundred Lectures with an Atlas of Eighty Plates, Representing Chemists, Institutions, Prime Materials, Crystals, Diagrams and Apparatus; and Illustrations in the Text** Gustavus Detlef Hinrichs, 1897      Lectures on Mineralogy T. Egleston, 2023-03-09 Reprint of the original first published in 1872

*Lectures on chemical pathology in its relation to practical medicine* Christian Archibald Herter, 1902      **Elementary Studies in Chemistry** Joseph Torrey (jr.), 1899      **Lectures on diseases of the digestive organs v. 1 1891** Carl Anton Ewald, 1891      *Register* Cornell University, 1905      Lectures on Agricultural Chemistry and Geology ... Appendix, containing suggestions for experiments in practical agriculture, etc James Finlay Weir Johnston, 1844      A Syllabus of a Course of Chemical Lectures Read at Guy's Hospital William Babington, William Allen, 1816      **Lectures on the Principle of Symmetry and Its Applications in All Natural Sciences** Francis Mauritius Jaeger, 1920      *Chemistry for Degree Students B.Sc. Semester - I (As per CBCS)* Madan R.L., This textbook has been designed to meet the needs of B Sc First Semester students of Chemistry as per the new UGC Model Curriculum Choice Based Credit System CBCS With its traditional approach to the subject this textbook lucidly explains principles of chemistry Important topics such as atomic structure chemical bonding molecular structure fundamentals of organic chemistry stereochemistry and aliphatic hydrocarbons are aptly discussed to give an overview of inorganic and organic chemistry Laboratory work has also been included to help students achieve solid conceptual understanding and learn experimental procedures

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