

Unit 18: Hydrocarbon and its Compounds

Things to be Studied in this Unit

- Introduction to hydrocarbon
- Saturated and unsaturated hydrocarbon
- Alkane, alkene and alkyne
- Homologous series
- Alkyl radical
- Functional group
- IUPAC system
- Common system to nomenclature of hydrocarbons
- Some hydrocarbon and its compounds
 - Methane
 - Propane
 - Alcohol
 - Ethane
 - Butane
 - Glycerol

Introduction

Hydrocarbons are organic compounds composed of hydrogen (H) and carbon (C) atoms. They are of two main types, viz. saturated and Unsaturated. The hydrocarbons in which carbon atoms are combined with a single bond are called saturated hydrocarbons. The hydrocarbons which contain a double bond or triple bond between carbon atoms are called unsaturated hydrocarbons. Hydrocarbons are the basis for many organic compounds and are found in various forms in fossil fuels like mineral oil and natural gas. They serve as essential building blocks for a wide range of chemical products and fuels, playing a vital role in industry and energy production.

Things to Know

-  **Hydrocarbon:** Hydrocarbons are the compounds formed by the combination of hydrogen and carbon. Examples: Methane (CH_4), Propane (C_3H_8) etc.
-  **Unsaturated hydrocarbon:** The hydrocarbon which contains a double bond or triple bond between carbon atoms is called unsaturated hydrocarbon. The general formula of unsaturated hydrocarbons is C_nH_{2n} (alkene) and $\text{C}_n\text{H}_{2n-2}$ (alkyne).
-  **Saturated hydrocarbon:** The hydrocarbon in which carbon atoms are combined with a single bond is called saturated hydrocarbon. The general formula of saturated hydrocarbon is $\text{C}_n\text{H}_{2n+2}$.
-  **IUPAC:** IUPAC stands for International Union of Pure and Applied Chemistry.
-  **Alkane:** An alkane is a saturated hydrocarbon in which carbon atoms are connected by only single covalent bonds. The general formula of alkanes is $\text{C}_n\text{H}_{2n+2}$ where 'n' is the number of carbon atoms present in one molecule of alkane. Example: Propane, Pentane, etc.
-  **Alkene:** An alkene is a saturated hydrocarbon in which carbon atoms are connected by a double bond. So an alkene contains $-\overset{\textstyle |}{\underset{\textstyle |}{\text{C}}}=\overset{\textstyle |}{\underset{\textstyle |}{\text{C}}}-$ group. The general formula of an alkene is C_nH_{2n} where 'n' is the number of carbon atoms in its one molecule. Example: Propene, Pentene, etc.

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Hazmat Chemistry Study Guide (Second Edition) Jill Meryl Levy, 2011

Student Study Guide/Solutions Manual for Essentials of General, Organic, and Biochemistry Denise Guinn, Rebecca Brewer, Rachel C. Lum, 2009-09-15 The Student Study Guide and Solutions Manual provides students with a combined manual designed to help them avoid common mistakes and understand key concepts After a brief review of each section s critical ideas students are taken through stepped out worked examples try it yourself examples and chapter quizzes all structured to reinforce chapter objectives and build problem solving techniques The solutions manual includes detailed solutions to all odd numbered exercises in the text
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