

(ix) The standard potential of the cell in the following reaction is _____



($E_{\text{Cd}^{2+}}^{\circ} = -0.403 \text{ V}$, $E_{\text{Cu}^{2+}}^{\circ} = 0.334 \text{ V}$)

(a) -0.737 V

(b) 0.737 V

(c) -0.069 V

(d) 0.069 V

(x) The value of $[\text{H}_3\text{O}^+]$ in mol lit^{-1} of 0.001 M acetic acid solution ($K_a = 1.8 \times 10^{-5}$) is _____

(a) 1.34×10^{-1}

(b) 1.34×10^{-2}

(c) 1.34×10^{-3}

(d) 1.34×10^{-4}

Q.2. Answer the following questions:

[8]

(i) Write the product formed when alkyl halide reacts with silver nitrite.

(ii) Write the name of product formed, when acetone is treated with 2, 4-dinitrophenyl hydrazine.

(iii) Write the name of biodegradable polyamide copolymer.

(iv) Identify the molecularity of following elementary reaction:



(v) What is the action of selenium on magnesium metal?

(vi) Write the name of isomerism in the following complexes:



(vii) Write the name of the alloy used in Fischer Tropsch process in the synthesis of gasoline.

(viii) Henry's law constant for $\text{CH}_3\text{Br}_{(\text{g})}$ is $0.159 \text{ mol dm}^{-3} \text{ bar}^{-1}$ at 25°C . What is solubility of $\text{CH}_3\text{Br}_{(\text{g})}$ in water at same temperature and partial pressure of 0.164 bar ?

SECTION – B

Attempt any EIGHT of the following questions:

[16]

Q.3. Explain pseudo-first order reaction with suitable example.

Q.4. Write the consequences of Schottky defect with reasons.

Q.5. What is the action of following on ethyl bromide:

(i) Na in dry ether

(ii) Mg in dry ether

Q.6. Explain formation of peptide linkage in protein with an example.

Q.7. Derive an expression to calculate molar mass of non volatile solute by osmotic pressure measurement.

Q.8. Explain monodentate and ambidentate ligands with example.

Q.9. Explain the trends in the following atomic properties of group 16 elements:

(i) Atomic radii

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Karin Nielsen-Saines



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1. Understanding the eBook Hsc Board Question Papers Science 2012
 - The Rise of Digital Reading Hsc Board Question Papers Science 2012
 - Advantages of eBooks Over Traditional Books
2. Identifying Hsc Board Question Papers Science 2012
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Hsc Board Question Papers Science 2012
 - User-Friendly Interface
4. Exploring eBook Recommendations from Hsc Board Question Papers Science 2012
 - Personalized Recommendations
 - Hsc Board Question Papers Science 2012 User Reviews and Ratings
 - Hsc Board Question Papers Science 2012 and Bestseller Lists

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 - Hsc Board Question Papers Science 2012 Budget-Friendly Options
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 - Hsc Board Question Papers Science 2012 Enhanced eBook Features
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 - Highlighting and Note-Taking Hsc Board Question Papers Science 2012
 - Interactive Elements Hsc Board Question Papers Science 2012
8. Staying Engaged with Hsc Board Question Papers Science 2012
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Hsc Board Question Papers Science 2012
9. Balancing eBooks and Physical Books Hsc Board Question Papers Science 2012
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Hsc Board Question Papers Science 2012
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Hsc Board Question Papers Science 2012
 - Setting Reading Goals Hsc Board Question Papers Science 2012
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Hsc Board Question Papers Science 2012
 - Fact-Checking eBook Content of Hsc Board Question Papers Science 2012
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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