

**Test for Anions :**

Anion	Test	Result
<b>Carbonate (<math>\text{CO}_3^{2-}</math>)</b>	Add dilute acid (e.g. dilute hydrochloric acid)	Effervescence, $\text{CO}_2$ produced (turns limewater milky)
<b>Sulfite (<math>\text{SO}_3^{2-}</math>)</b>	Add dilute acid (e.g. dilute hydrochloric acid), warm	$\text{SO}_2$ produced (turns aq. potassium manganate (VII) from purple to colourless)
<b>Chloride (<math>\text{Cl}^-</math>)</b> [in solution]	Acidify, then add aq. silver nitrate [ $\text{Ag}^+$ ] Acidify, then add aq. lead nitrate [ $\text{Pb}^{2+}$ ]	White ppt. of $\text{AgCl}$ $\text{Ag}^+ (\text{aq}) + \text{Cl}^- (\text{aq}) \rightarrow \text{AgCl} (\text{s})$ White ppt. of $\text{PbCl}_2$
<b>Bromide (<math>\text{Br}^-</math>)</b> [in solution]	Acidify, then add aq. silver nitrate	Cream ppt. of $\text{AgBr}$
<b>Iodide (<math>\text{I}^-</math>)</b> [in solution]	Acidify, then add aq. silver nitrate (or lead nitrate)	Yellow ppt. of $\text{AgI}$ (or $\text{PbI}_2$ )
<b>Sulfate (<math>\text{SO}_4^{2-}</math>)</b> [in solution]	Acidify, then add aq. barium nitrate (or barium chloride) [ $\text{Ba}^{2+}$ ]	White ppt. of $\text{BaSO}_4$ $\text{Ba}^{2+} + \text{SO}_4^{2-} \rightarrow \text{BaSO}_4$
<b>Nitrate (<math>\text{NO}_3^-</math>)</b>	Add aq. sodium hydroxide, then little aluminium (foil) and warm carefully	Ammonia produced (turns damp red litmus paper or universal indicator paper blue)

# June 2002 Paper 6 Chemistry Igcse

**Zhe Xu**



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