

## Equilibrium Practice Problems

1. Write the equilibrium expression for each of the following reactions:



$$K = \frac{[\text{NH}_3]^2}{[\text{N}_2] [\text{H}_2]^3}$$



$$K = \frac{[\text{ICl}]^2}{[\text{Cl}_2]}$$



$$K = \frac{[\text{NO}] [\text{O}_2]^{1/2}}{[\text{NO}_2]}$$

2. The dissociation of acetic acid,  $\text{CH}_3\text{COOH}$ , has an equilibrium constant at  $25^\circ\text{C}$  of  $1.8 \times 10^{-5}$ . The reaction is



If the equilibrium concentration of  $\text{CH}_3\text{COOH}$  is 0.46 moles in 0.500 L of water and that of  $\text{CH}_3\text{COO}^-$  is  $8.1 \times 10^{-3}$  moles in the same 0.500 L, calculate  $[\text{H}^+]$  for the reaction.

$$K = \frac{[\text{CH}_3\text{COO}^-] [\text{H}^+]}{[\text{CH}_3\text{COOH}]} \quad \text{Plug in known values and solve.}$$

$$1.8 \times 10^{-5} = \frac{[8.1 \times 10^{-3} \text{ moles}/0.500 \text{ L}] [\text{H}^+]}{[0.46 \text{ moles}/0.500 \text{ L}]}$$

$$[\text{H}^+] = 1.0 \times 10^{-3} \text{ M}$$

3. Indicate the effect of a catalyst, pressure, temperature and concentration on each of the following on:

	catalyst	pressure	temperature	concentration
a. speed	presence increase	direct relationship	direct relationship	direct relationship
b. pos <sup>n</sup>	nothing	based on moles	based on enthalpy of rxn	inverse relationship

# Honors Chemistry Equilibrium Practice

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## **Honors Chemistry Equilibrium Practice:**

**Equilibrium Problems and Applications** Gábor Kassay, Vicențiu Rădulescu, 2018-10-09 Equilibrium Problems and Applications develops a unified variational approach to deal with single valued set valued and quasi equilibrium problems The authors promote original results in relationship with classical contributions to the field of equilibrium problems The content evolved in the general setting of topological vector spaces and it lies at the interplay between pure and applied nonlinear analysis mathematical economics and mathematical physics This abstract approach is based on tools from various fields including set valued analysis variational and hemivariational inequalities fixed point theory and optimization Applications include models from mathematical economics Nash equilibrium of non cooperative games and Browder variational inclusions The content is self contained and the book is mainly addressed to researchers in mathematics economics and mathematical physics as well as to graduate students in applied nonlinear analysis A rigorous mathematical analysis of Nash equilibrium type problems which play a central role to describe network traffic models competition games or problems arising in experimental economics Develops generic models relevant to mathematical economics and quantitative modeling of game theory aiding economists to understand vital material without having to wade through complex proofs Reveals a number of surprising interactions among various equilibria topics enabling readers to identify a common and unified approach to analysing problem sets Illustrates the deep features shared by several types of nonlinear problems encouraging readers to develop further this unifying approach from other viewpoints into economic models in turn

**Honors in Practice** ,2009 [AP Chemistry For Dummies](#) Peter J. Mikulecky, Michelle Rose Gilman, Kate Brutlag, 2008-11-13 A practical and hands on guide for learning the practical science of AP chemistry and preparing for the AP chem exam Gearing up for the AP Chemistry exam AP Chemistry For Dummies is packed with all the resources and help you need to do your very best Focused on the chemistry concepts and problems the College Board wants you to know this AP Chemistry study guide gives you winning test taking tips multiple choice strategies and topic guidelines as well as great advice on optimizing your study time and hitting the top of your game on test day This user friendly guide helps you prepare without perspiration by developing a pre test plan organizing your study time and getting the most out of your AP course You ll get help understanding atomic structure and bonding grasping atomic geometry understanding how colliding particles produce states and so much more To provide students with hands on experience AP chemistry courses include extensive labwork as part of the standard curriculum This is why the book dedicates a chapter to providing a brief review of common laboratory equipment and techniques and another to a complete survey of recommended AP chemistry experiments Two full length practice exams help you build your confidence get comfortable with test formats identify your strengths and weaknesses and focus your studies You ll discover how to Create and follow a pretest plan Understand everything you must know about the exam Develop a multiple choice strategy Figure out displacement combustion and acid base reactions Get

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