

## Honors Chemistry Rate Law Homework

- 1) For the reaction  $A + B \rightarrow C$ , we've determined the following experimental data:

Time	$[A]$ (M)	$[B]$ (M)	Initial rate (M/s)
1	0.0100	0.0100	$1.82 \times 10^{-3}$
2	0.0200	0.0100	$3.24 \times 10^{-3}$
3	0.0100	0.0200	$4.48 \times 10^{-3}$

Using this information, determine the rate law for this reaction, find the reaction order for A and B, and find the overall reaction order.

- 2) In the reaction  $A + B \rightarrow C$ , we found the following rate data:

Time	$[A]$ (M)	$[B]$ (M)	Initial rate (M/s)
1	0.005	0.015	$2.85 \times 10^{-3}$
2	0.005	0.030	$5.69 \times 10^{-3}$
3	0.010	0.015	$11.3 \times 10^{-3}$

Using this information, find the overall rate law for this reaction, find the order of the reaction for each reactant, and the overall reaction order.

- 3) Using the information from the rate equations you found in problem 2, determine the rate constant  $k$  for this reaction.

# Honors Chemistry Rate Law Homework

**Tao Wei**



## Honors Chemistry Rate Law Homework:

*University Bulletin* University of California (System), 1965      *Martindale-Hubbell International Law Directory* ,1993  
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- Chapter 1: Introduction to Honors Chemistry Rate Law Homework
- Chapter 2: Essential Elements of Honors Chemistry Rate Law Homework
- Chapter 3: Honors Chemistry Rate Law Homework in Everyday Life
- Chapter 4: Honors Chemistry Rate Law Homework in Specific Contexts
- Chapter 5: Conclusion

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6. In chapter 5, the author will draw a conclusion about Honors Chemistry Rate Law Homework. The final chapter will summarize the key points that have been discussed throughout the book.

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