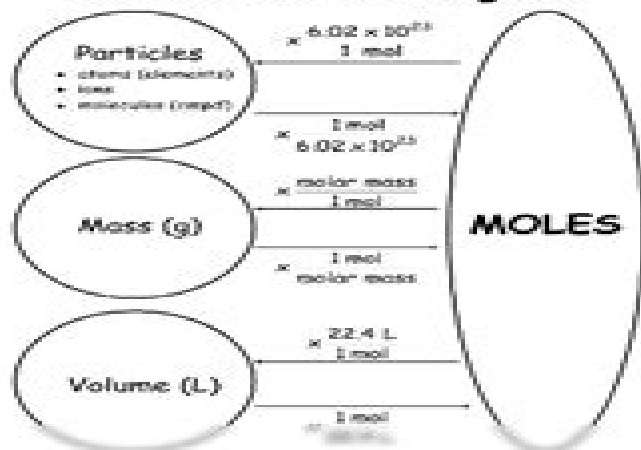


Honors Chemistry B- Final Exam Formula Sheet

Note: You will not use all of these formulas for this test

Temperature Formulas:
 $^{\circ}\text{C} = (^{\circ}\text{F} - 32) \cdot \frac{5}{9}$ $^{\circ}\text{F} = \frac{9}{5}^{\circ}\text{C} + 32$ $\text{K} = ^{\circ}\text{C} + 273$

Mole Conversion Diagram



1 atm = 760 mm Hg = 101.3 kPa

$P_1 + P_2 = P_3 + P_4$

$\frac{m_1}{m_2} = \frac{d_1}{d_2}$

$\frac{m_1}{V_1} = \frac{m_2}{V_2}$

$\Delta \frac{m}{V} = \frac{\Delta m}{V} = \frac{\Delta m}{V}$

$P_1 + P_2 + P_3 + P_4 = P_5 + P_6$

$P_{\text{total}} = P_1 + P_2 + P_3 + \dots$

$\frac{m_1}{V_1} = \frac{m_2}{V_2}$

$\frac{m_1}{V_1} = \frac{m_2}{V_2} = \frac{m_3}{V_3} = \dots$

$M_1 \times V_1 = M_2 \times V_2$

Percent by Mass = $\frac{\text{Mass of Solute}}{\text{Mass of Solution}} \times 100$

Percent by Volume = $\frac{\text{Volume of Solute}}{\text{Volume of Solution}} \times 100$

Mole fractions: $X_A = \frac{n_A}{n_A + n_B}$ $X_B = \frac{n_B}{n_A + n_B}$

1 Calorie = 1 kilocalorie = 1000 calories

$Q = mc\Delta T$

$q_{\text{sys}} = \Delta H = -q_{\text{surr}} = -m \times C \times \Delta T$

$\Delta H^{\circ} = \Delta H_f^{\circ}(\text{products}) - \Delta H_f^{\circ}(\text{reactants})$

$K_c = \frac{[\text{C}]^c \cdot [\text{D}]^d}{[\text{A}]^a \cdot [\text{B}]^b}$

$\Delta G = \Delta H - T\Delta S$

$\ln K = -\frac{\Delta G^{\circ}}{RT}$

$K_w = [\text{H}^+] \cdot [\text{OH}^-] = 1.0 \times 10^{-14}$

$\text{pH} = -\log[\text{H}^+]$

$\text{pOH} = -\log[\text{OH}^-]$

$K_a = \frac{[\text{H}^+][\text{A}^-]}{[\text{HA}]}$

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Michael Rajnik



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