

- 3.5.** In a homogeneous isothermal liquid polymerization, 20% of the monomer disappears in 34 minutes for initial monomer concentration of 0.04 and also for 0.8 mol/liter. What rate equation represents the disappearance of the monomer?

Since the fractional disappearance is independent of initial concentration we have first order rate equation,

$$\ln\left(\frac{C_{a0}}{C_a}\right) = Kt$$

Where, C_a = Monomer concentration

We also can find the rate constant, thus replacing values.

Case I: $C_{a0} = 0.04$ mol/l

$$\ln\left(\frac{0.04}{0.8+0.04}\right) = k*34$$

So, $k = 0.00657 \text{ min}^{-1}$.

Case II: $C_{a0} = 0.8$ mol/l

$$\ln\left(\frac{0.8}{0.8+0.8}\right) = k*34$$

$k = 0.00657 \text{ min}^{-1}$.

Hence the rate of disappearance of monomer is given by,

$$-r_a = (0.00657 \text{ min}^{-1}) * C$$

- 3.6.** After 8 minutes in a batch reactor, reactant ($C_{A0} = 1$ mol/liter) is 80% converted; after 18 minutes, conversion is 90%. Find a rate equation to represent this reaction.

$$C_{a0} = 1 \frac{\text{mol}}{\text{lit}}$$

$t = 8$ min then conversion 80%

$t = 18$ mins then conversion 90%

Assume the reaction is second order

Levenspiel Chapter 12 Solution Manual

Camilla Rothe



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