

1. What volume of hydrogen gas at STP is produced from the reaction of 50.0g of Mg and 75.0 grams of HCl? How much of the excess reagent is left over (in grams)?



Do a standard mass to mass problem starting with each reactant

50.0 g Mg	1 mol Mg 24.31 g Mg	1 mol H ₂ 1 mol Mg	22.4 L 1 mol H ₂	= 46.1 L H ₂
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75.0 g HCl	1 mol HCl 36.46 g HCl	1 mole H ₂ 2 mol HCl	22.4 L 1 mol H ₂	= 23.0 L H ₂
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HCl is the limiting reactant!!

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